

Titration and Neutralisation

Chemistry

Week 10

Name: _____

Date: _____

Acid-Base Titration

An acid-base **titration** is a technique by which the concentration of an acid or base is determined by reacting it with a known volume of a base or acid of known concentration. The **equivalence point** is the point at which the acid and base have reacted in a 1:1 molar ratio. The **end point** is the point at which the indicator changes color. The **equivalence point** is the point at which the acid and base have reacted in a 1:1 molar ratio. The **end point** is the point at which the indicator changes color.

$$n(A^-) = n(H^+)$$

Equation

The concentration of acid

$n(A^-)$ is moles of acid

$n(H^+)$ is moles of base

$n(A^-)$ is moles of acid

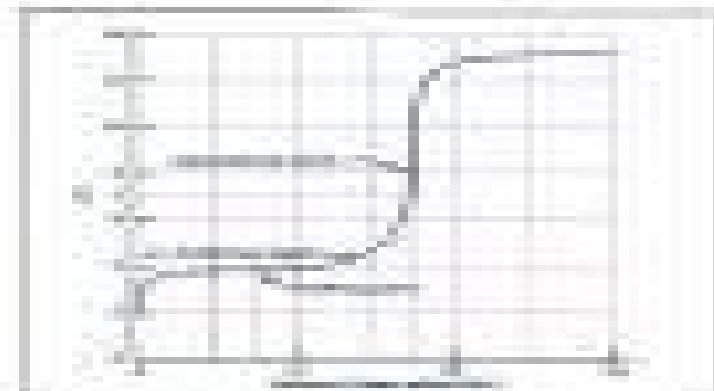
When performing acid-base titrations, an indicator is used to indicate when the acid and base have reacted in a 1:1 molar ratio. The indicator is a substance that changes color when the **equivalence point** is reached. The indicator is chosen based on the **pH range** of the equivalence point. The **equivalence point** is the point at which the acid and base have reacted in a 1:1 molar ratio. The **end point** is the point at which the indicator changes color. The **equivalence point** is the point at which the acid and base have reacted in a 1:1 molar ratio. The **end point** is the point at which the indicator changes color.

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Equivalence point: $n(A^-) = n(H^+)$

Titration And Neutralization Chemistry Week 34

Gordon Mabey Kline



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