

Unit 9 Outline- Thermodynamics, Reaction Rates, & Chemical Equilibrium

Chapters 17 & 18

Essential Skills/ State Standards:

- 9-1. Describe how temperature and heat flow in terms of the motion of molecules (or atoms). (CSS: 7a)
 9-2. Know that chemical processes & phase changes can either release energy (exothermic) or absorb energy (endothermic). (CSS: 7b, 7c)
 9-3. Solve problems involving heat flow and temperature changes, using known values of specific heat ($q = m \cdot c \cdot \Delta T$) & latent heat of phase change ($q = m \cdot \Delta H$). (CSS: 7d)
 9-4. Know the rate of reaction is the decrease in concentration of reactants or the increase in concentration of products with time & depends on such factors as: concentration, temperature, and pressure. (CSS: 8a, 8b)
 9-5. Know the definition and role of activation energy & catalyst in a chemical reaction. (CSS: 8d*, 8c)
 9-6. Know how to use Le Chatelier's principle to predict the effect of changes in concentration, temperature, and pressure (CSS: 9a)
 9-7. Know that equilibrium is established when the forward and reverse reaction rates are equal. (CSS: 9b)

Vocabulary

- | | | | |
|---|---|--|--|
| <ul style="list-style-type: none"> • Condensation • Freezing • Vaporization • Melting • Boiling • Sublimation | <ul style="list-style-type: none"> • Deposition • Heat • Temperature • Thermochemistry • Exothermic • Endothermic | <ul style="list-style-type: none"> • Calorimeter • Joule • Specific heat • $q = m \cdot c \cdot \Delta T$ • $q = m \cdot \Delta H_{\text{vaporization}}$ • $q = m \cdot \Delta H_{\text{fusion}}$ | <ul style="list-style-type: none"> • Heat of reaction • Le Chatelier's principle • Equilibrium • Activation Energy • Catalyst |
|---|---|--|--|

Book Assignments:

Topic	Read	Required Problems
1. Heat, Temperature & Specific Heat	Section 17-1 p. 511-524	Section 17-p. 524 #2, 5 & 6 Chapter 17 Review p. 548 #16 & 17
3. Chemical Equilibrium & Le Chatelier's Principle	Section 12-3 p. 372-375 Section 18-2 p. 562-568	Chapter 18 Review p. 586 #4, 6 & 7
4. Phase Changes Diagrams & Latent Heat of Phase Change	Section 12-3 p. 378-382	Section 12-3 p. 382 #4-8 Chapter 12 Review p. 389 #19, 28, 31 & 37
5. Collision Theory & Activation Energy	Section 17-3 p. 532-534	Section 17-3 p. 537 #2, 3 & 5
6. Reaction Rates & Catalysts	Section 17-4 p. 538-544	Section 17-4 p. 545 #2-4

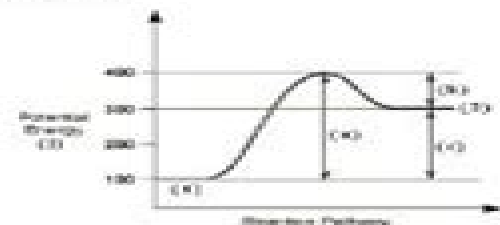
Study Guide (Use notes, worksheets, and labs in addition to this study guide)

- According to the collision theory, what 3 things must happen in order for two molecules to react?
- What is the importance of the activation energy in a chemical reaction?
- Using the collision theory, explain if the following would increase/ decrease the rate of reaction:

$$2 \text{HCl(aq)} + \text{Mg(s)} \rightarrow \text{MgCl}_2\text{(aq)} + \text{H}_2\text{(g)}$$

a. Increase temp:	b. Increase HCl	c. Increase pressure	d. Remove MgCl_2	e. Add H_2
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- Identify the following terms with their correct number/letter from the graphs:

• Reactants-	• Products-
• Activation energy (number/letter & amount in Joules)-	
• Is this reaction exothermic or endothermic?	
- Explain how a catalyst affects the activation energy in a chemical reaction. (Also, show on the above graph what a catalyst might do).



Unit 7chemistry Stoichiometry Study Guide Answers

Margaret G. Kimble



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Cambridge IGCSE™ Chemistry Teacher's Guide (Collins Cambridge IGCSE™) Chris Sunley,2022-02-03 Prepare students with complete coverage of the revised Cambridge IGCSE™ Chemistry syllabus 0620 0971 for examination from 2023 Collins Cambridge IGCSE Chemistry Teacher's Guide is full of lesson ideas practical instructions technician's notes planning support and more

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Experts,2020-06-20 **Student Study Guide Fundamentals of Chemistry by David E. Goldberg** Margaret G.

Kimble,1994 Self-Learner to ICSE Chemistry class 10 Baljinder kaur, Self Learner to ICSE Chemistry Class 10 is a comprehensive solution book crafted for self study and board exam mastery It provides step by step explanations to all questions from the prescribed ICSE Chemistry textbook The chapters begin with concise summaries formulas and definitions making complex topics easier to grasp Each exercise includes structured reasoning based and numerical questions presented with clear examiner style answers Designed for independent learners it helps build a deep understanding of concepts such as atomic structure chemical bonding acids and bases and organic chemistry The Self Learner series promotes conceptual clarity confidence and thorough revision With ample solved examples practice tests and additional board style questions this book ensures students are exam ready and capable of achieving outstanding scores in ICSE Chemistry

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Table of Contents Unit 7chemistry Stoichiometry Study Guide Answers

1. Understanding the eBook Unit 7chemistry Stoichiometry Study Guide Answers
 - The Rise of Digital Reading Unit 7chemistry Stoichiometry Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Unit 7chemistry Stoichiometry Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Unit 7chemistry Stoichiometry Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Unit 7chemistry Stoichiometry Study Guide Answers
 - Personalized Recommendations
 - Unit 7chemistry Stoichiometry Study Guide Answers User Reviews and Ratings

- Unit 7chemistry Stoichiometry Study Guide Answers and Bestseller Lists
- 5. Accessing Unit 7chemistry Stoichiometry Study Guide Answers Free and Paid eBooks
 - Unit 7chemistry Stoichiometry Study Guide Answers Public Domain eBooks
 - Unit 7chemistry Stoichiometry Study Guide Answers eBook Subscription Services
 - Unit 7chemistry Stoichiometry Study Guide Answers Budget-Friendly Options
- 6. Navigating Unit 7chemistry Stoichiometry Study Guide Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Unit 7chemistry Stoichiometry Study Guide Answers Compatibility with Devices
 - Unit 7chemistry Stoichiometry Study Guide Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Unit 7chemistry Stoichiometry Study Guide Answers
 - Highlighting and Note-Taking Unit 7chemistry Stoichiometry Study Guide Answers
 - Interactive Elements Unit 7chemistry Stoichiometry Study Guide Answers
- 8. Staying Engaged with Unit 7chemistry Stoichiometry Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Unit 7chemistry Stoichiometry Study Guide Answers
- 9. Balancing eBooks and Physical Books Unit 7chemistry Stoichiometry Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Unit 7chemistry Stoichiometry Study Guide Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Unit 7chemistry Stoichiometry Study Guide Answers
 - Setting Reading Goals Unit 7chemistry Stoichiometry Study Guide Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Unit 7chemistry Stoichiometry Study Guide Answers
 - Fact-Checking eBook Content of Unit 7chemistry Stoichiometry Study Guide Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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