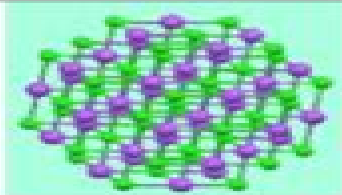
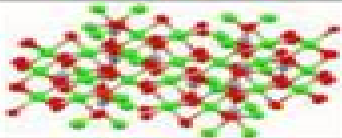
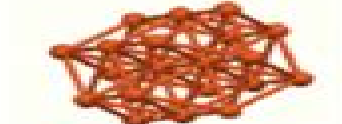


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In the table below, identify the type of solid (atomic, molecular, ionic) each substance forms. Then predict in which phase (solid, liquid or gas) each substance would exist on the Earth, on Mercury and on Pluto.

| Substance                     | Crystal structure                                                                   | Type     | melting pt (°C)<br>boiling pt (°C) | Phase at room temp<br>on                                                     |
|-------------------------------|-------------------------------------------------------------------------------------|----------|------------------------------------|------------------------------------------------------------------------------|
| NaCl                          |    | <b>I</b> | 801<br>1413                        | Earth (25°C) <b>S</b><br>Mercury (450°C) <b>S</b><br>Pluto (-230°C) <b>S</b> |
| CaCO <sub>3</sub>             |    | <b>I</b> | 520<br>decomposes at 825           | E <b>S</b><br>M <b>S</b><br>P <b>S</b>                                       |
| S <sub>8</sub>                |    | <b>M</b> | 115<br>444                         | E <b>S</b><br>M <b>G</b><br>P <b>S</b>                                       |
| Cu                            |   | <b>A</b> | 1084<br>2567                       | E <b>S</b><br>M <b>S</b><br>P <b>S</b>                                       |
| H <sub>2</sub> O <sub>2</sub> |  | <b>M</b> | -0.41<br>150                       | E <b>L</b><br>M <b>G</b><br>P <b>S</b>                                       |

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