

P3 – Relate chemical reactions of elements and their compounds to the position of the elements in the periodic table:

D1 – Analyse the extent to which properties of elements and their compounds may be predicted by knowing the position of the elements in the periodic table:

Period 2 elements react with oxygen:

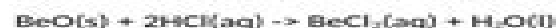
Lithium – it is arranged in a giant metallic lattice; it is surrounded by a lot of delocalised electrons. It is solid and it is less dense than water, this makes it a good conductor of electricity and heat. When lithium is burned it turns white and forms lithium oxide when reacts with oxygen.



When lithium reacts with water it easily dissolves in it forming an alkaline solution called the lithium hydroxide. The lithium hydroxide makes a strong base alkali, it is a soluble base too and therefore forms salt. The lithium oxide is only basic. It is not very reactive, the lower the alkali metal in the group the more reactive it becomes. Lithium is not as reactive as potassium due to the position of the element on the periodic table in the group 2.



Beryllium– it is arranged in a giant metallic lattice, just like lithium it is surrounded by delocalised electrons. It is usually in a solid form at room temperatures and due to its structure it has a very high melting and boiling point. It is capable of conducting electricity and heat, but it is not very good at it. When beryllium reacts with oxygen it forms a white powder called beryllium oxide. The bonds which are formed when heated strongly are intermediate ionic-covalent bonds. When water is added into the beryllium oxide, the mixture is insoluble which means that it doesn't dissolve in the water – therefore no reaction would be formed. The beryllium oxide on the other hand is able to react with acids due to being an amphoteric oxide. When it reacts with an acid e.g. beryllium oxide reacts with hydrochloric acids (one of the most common acids used in laboratories) it forms beryllium chloride.



The beryllium oxide is on the other hand a basic oxide which dissolves in acid in order to form different forms of salt. When beryllium oxide (BeO) is added to the sulphuric acid it forms beryllium sulphate and water as products.



It is able to dissolve in strong acids; it would eventually form beryllate(II) salts. It can be formed by adding beryllium oxide with sodium hydroxide, this would form sodium beryllate(II). This allows beryllium to be able to be transformed into different chemicals. When beryllium is being heated with chlorine, it forms a covalent bonding with chlorine resulting in beryllium chloride. Beryllium oxide is not very reactive; it is a lot less reactive than potassium. On the reactivity series, the beryllium itself is located below magnesium making it less reactive. Beryllium oxide doesn't react with water.

Unit 27 Chemical Periodicity Its Applications

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