

Titration and Neutralisation

Chemistry

Week 10

Name: _____

Date: _____

Acid-Base Titration

An acid-base **titration** is a technique by which the concentration of an acid or base is determined by reacting it with a known volume of a base or acid of known concentration. The **equivalence point** is the point at which the acid and base have reacted in a 1:1 ratio. The **end point** is the point at which the indicator changes color. The **equivalence point** is the point at which the acid and base have reacted in a 1:1 ratio. The **end point** is the point at which the indicator changes color.

$$n(A^-) = n(H^+)$$

Equation

The concentration of acid

$n(A^-)$ is moles of acid

$n(H^+)$ is moles of base

$n(A^-)$ is moles of acid

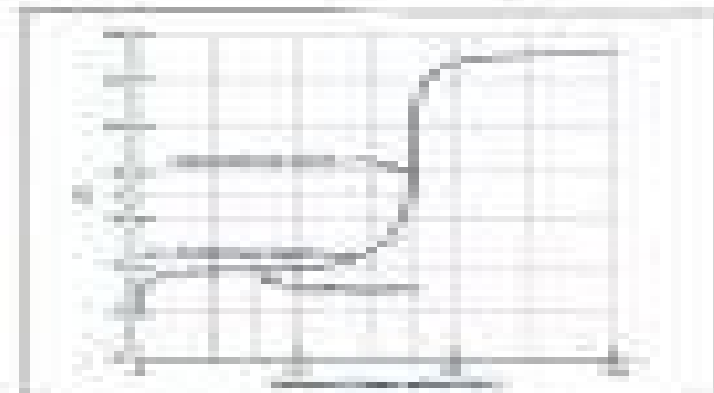
When performing acid-base titrations, an indicator is used to indicate when the acid and base have reacted in a 1:1 ratio. The indicator changes color when the **equivalence point** is reached. If the indicator is not chosen correctly, the end point will not be the same as the equivalence point. This is not always the case, but often it is. This point is known as the **equivalence point**. For strong acid-strong base titrations, the **equivalence point** occurs at pH 7.0.

When you add an acid to a base, the pH of the solution changes. The change in pH is called a **pH change**. A **pH change** occurs when the acid and base react. The **pH change** is the difference between the pH of the acid and the pH of the base. The **pH change** is the difference between the pH of the acid and the pH of the base. The **pH change** is the difference between the pH of the acid and the pH of the base.

The **equivalence point** is the point at which the acid and base have reacted in a 1:1 ratio. The **equivalence point** is the point at which the acid and base have reacted in a 1:1 ratio. The **equivalence point** is the point at which the acid and base have reacted in a 1:1 ratio. The **equivalence point** is the point at which the acid and base have reacted in a 1:1 ratio.

When the graph is plotted, the **equivalence point** is the point at which the acid and base have reacted in a 1:1 ratio. The **equivalence point** is the point at which the acid and base have reacted in a 1:1 ratio. The **equivalence point** is the point at which the acid and base have reacted in a 1:1 ratio.

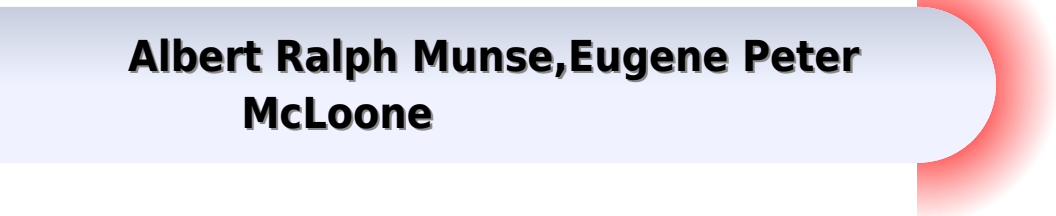
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Equivalence point occurs at pH 7.0

Titration And Neutralization Chemistry Week 34

**Albert Ralph Munse, Eugene Peter
McLoone**



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