

Chemistry 30 Thermochemistry Quiz #2 Practice

Use the following information to answer the next question:

The combustion of ethanol ($\text{C}_2\text{H}_5\text{OH}$) at 298 K, standard state ($\text{C}_2\text{H}_5\text{OH}(l)$), can be represented by the following thermochemical equation:



Equation:



$$\Delta H^\circ = -1366.8 \text{ kJ}$$



$$\Delta H^\circ = -455.6 \text{ kJ}$$

4. The enthalpy change for the combustion of ethanol as written in the thermochemical equation is:

- A. -1366.8 kJ
- B. -455.6 kJ
- C. -1366.8 kJ
- D. -455.6 kJ

Use the following information to answer the next question:

When carbon is completely combusted (combustion), it represents the following equation:



5. The maximum heat enthalpy information with the reaction is

- A. -393.5 kJ/mol
- B. -393.5 kJ/mol
- C. -393.5 kJ/mol
- D. -393.5 kJ/mol

Thermochemistry Question Answer Guide

Bernice G. Segal, Peter S. Shenkin



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