

Types of Reactions Lab



Name _____

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INTRODUCTION

Matter can undergo both physical and chemical changes. Chemical changes result in the formation of new materials. Observable signs of chemical reaction include the release of a gas, a change in color, the formation of a solid, and the production of heat and light.

Chemical reactions can be classified in many ways. One classification system involves five general types of reactions: synthesis, decomposition, single displacement, double displacement, and combustion. In this lab, you will perform several reactions and classify them as one or more of the five general types.

SAFETY

- When heating a test tube, aim the mouth of a test tube AWAY from yourself and others.
- Do NOT look directly at burning magnesium. The intense light may damage your eyes.

PROCEDURE

1. Create a table of observations that includes 5 columns: Reactant(s), Appearance of Reactant(s) Before Mixing, Mass of Reactant(s), Appearance of Product(s), Mass of Product(s).

Reaction A

2. Obtain 5 mL of sodium sulfate (Na_2SO_4) solution in a clean test tube. Next, obtain 5 mL of barium chloride (BaCl_2) solution in another test tube. Place both test tubes in a beaker and measure the mass of the beaker with its contents. Record your observations. Pour the contents from one test tube into the other. Place both test tubes back into the beaker and measure the mass of the beaker once again. Record your observations.

Reaction B

3. Obtain 5 mL of copper(II) sulfate solution in a clean test tube. Obtain a small piece of steel wool (iron). Place the steel wool into a beaker, and the test tube into the same beaker. Measure the mass of the beaker with its contents. Record your observations. Place the steel wool into the test tube and observe the reaction. Place the test tube back into the beaker and measure the mass of the beaker once again. Record your observations.

Reaction C

4. Obtain 5 mL of hydrogen peroxide (H_2O_2) in a clean test tube. Use a scoopula to obtain an extremely small amount of the catalyst manganese(IV) oxide (MnO_2), the size of a match head, and place it in a beaker. Place the test tube into the same beaker and measure the mass of the beaker with its contents. Record your observations. Pour the manganese(IV) oxide into the test tube and observe the reaction. After a few seconds, insert a glowing splint into the mouth of the test tube upon which the splint will burst into flames. Place the test tube back into the beaker and measure the mass of the beaker once again. Record your observations.

Reaction D

5. Obtain a piece of magnesium ribbon approximately 1 cm long in a watch glass. Measure the mass of the watch glass with its contents. Record your observations. Using crucible tongs, hold the ribbon in the burner flame until the magnesium starts to burn. DO NOT LOOK DIRECTLY AT THE FLAME. Hold the burning magnesium directly over the watch glass. When the ribbon stops burning, place the remains on the watch glass. Measure the mass of the watch glass once again. Record your observations. Turn off the burner.
6. Clean your test tubes and lab station. Wash liquids down the drain. Wrap solids in a paper towel before discarding them in the trashcan.

Types Of Reactions Lab

**Howard Crosby Warren, Clement Leslie
Vaughan, Knight Dunlap, Madison
Bentley, Walker Samuel Hunter**

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A Concrete Stoichiometry Unit for High School Chemistry Jennifer Louise Pakkala, 2006 **Psychological Index**, 1896 **The Psychological Index**, 1895 **The Psychological Review. Monograph Supplement**, 1896

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