



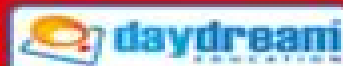
The Pocket-Sized Revision Guide

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WJEC

Higher Tier

For Maths & Maths Numeracy



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2012 GCSE English results Great Britain: Parliament: House of Commons: Education Committee, 2013-06-11 The 2012 GCSE English results prompted significant controversy which ultimately resulted in an application for judicial review This report sets out the background to these events and identifies lessons to be learned The problems with GCSE English can be traced back to the 2007/09 development phase of the qualification in particular the turbulence which resulted from the shift away from a mostly linear to a modular system combined with a high proportion of controlled assessment and generous marking tolerances Exam board experts raised concerns at the time but these were not acted upon by the regulator the then interim Ofqual Further difficulties arose because of pressures from the school accountability system The problems experienced with GCSE English in 2012 highlighted serious weaknesses in the moderation of speaking and listening with consequences for grade awarding The current status of Ofqual as an independent regulator accountable to Parliament is the right one However the Coalition Government is bringing in wholesale changes to GCSEs and A levels to a tight timetable and at the same time Ofqual must have systems in place The Committee is also concerned that there is a rush towards separate exam systems for England Wales and Northern Ireland without careful reflection on what might be lost or consensus that this is the right thing to do New WJEC GCSE Maths Exam Practice Workbook: Higher (includes Answers) CGP

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web the equation $\text{mg hcl} + \text{mgcl}_2 + \text{h}_2$ is clearly unbalanced because on one side there is only 1 hydrogen atom but on the other side there are 2 also unbalanced chlorine but they both come from the same compound so this becomes really simple to

balance this equation is easily balanced by placing the coefficient 2 in front of molecule

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there are equal numbers of atoms of each element on both sides of the equation we can write a chemical equation for the reaction of carbon with hydrogen gas to form methane CH_4

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$$3 \text{ C}_2\text{H}_2 + 2 \text{ KClO}_3 \rightarrow 2 \text{ C}_2\text{H}_4 + 2 \text{ KCl} + 3 \text{ O}_2$$

$$2 \text{ NaCl} + 1 \text{ F}_2 \rightarrow 2 \text{ NaF}$$

$$1 \text{ Cl}_2 + 4 \text{ H}_2 \rightarrow 2 \text{ H}_2\text{O} + 5 \text{ H}_2$$

$$1 \text{ Pb} + \text{OH}^- + 2 \text{ HCl} \rightarrow 2 \text{ H}_2\text{O} + 1 \text{ PbCl}_2$$

$$2 \text{ Al} + 3 \text{ K}_2\text{SO}_4 + 6 \text{ KBr} \rightarrow 1 \text{ Al}_2\text{SO}_4 + 3 \text{ K}_2\text{SO}_4 + 7 \text{ H}_2$$

$$1 \text{ CH}_4 + 2 \text{ O}_2 \rightarrow 1 \text{ CO}_2 + 2 \text{ H}_2\text{O}$$

$$1 \text{ C}_3\text{H}_8 + 5 \text{ O}_2 \rightarrow 3 \text{ CO}_2 + 4 \text{ H}_2\text{O}$$

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