

LCD Color Television

32AV623D

Ver. 2.01

This model is classified as a green product (*1), as indicated by the underlined serial number. This Service Manual describes replacement parts for the green product. When repairing this green product, use the part(s) described in this manual and lead-free solder (*2). For (*1) and (*2), refer to **GREEN PRODUCT PROCUREMENT** and **LEAD-FREE SOLDER**.

Toshiba 32av623d Lcd Tv Service Manual

S. Prentiss



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TV Service Manual S. Prentiss,1973-01-01 *Japanese Color Tv Service Manual. No. 1. Sony, Sharp, Midland S.*
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Service Manual Vols. 1 Stan Prentiss,1970 *Japanese Color TV Service Manual No. 1* Stan Prentiss,1971 **General**
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Books,1975-01-01 **Japanese Color Tv Service Manual. No. 2. Panasonic, Hitachi, Delmonico** S. Prentiss,1971
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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) + \text{C}_2\text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_4(\text{g}) + \text{O}_2(\text{g}) + \text{C}_2\text{H}_6(\text{g})$ $\Delta h = 94.5 \text{ kJ}$
 $\text{H}_2\text{O}(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$ $\Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{g})$ $\Delta h = 283 \text{ kJ}$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{l})$ $\Delta h = 1411 \text{ kJ/mole}$
 $\text{C}_2\text{H}_6(\text{g}) + 7/2 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{l})$ $\Delta h = 1560 \text{ kJ/mole}$
 $\text{H}_2(\text{g}) + 1/2 \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$ $\Delta h = 285 \text{ kJ/mole}$

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 hess s law example 12.1 determine Δh for the target reaction
 $\text{C}_2\text{H}_4(\text{g}) + \text{NO}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g}) + \text{O}_2(\text{g}) + \text{N}_2\text{O}_5(\text{g})$
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 $\text{Mg}(\text{s}) + \text{O}_2(\text{g}) \rightarrow \text{MgO}(\text{s})$ we used

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 cohen 2016 the equations for the reactions in the experiment done are as follows
 $\text{NaOH}(\text{s}) + \text{Na}(\text{aq}) + \text{OH}(\text{aq})$

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web thus for reaction 9 2 in which 1 mole of H_2SO_4 and 1 mole of OH^- combine to form 2 moles of H_2O $\Delta h = 55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta h_{\text{products}} - \Delta h_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta h_{\text{products}} > \Delta h_{\text{reactants}}$

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