

Directions:

- You must bring a pencil and calculator to the exam.
- You will be given a periodic table.
- If you complete the entire review sheet by the day of your exam you will be eligible for the curve should one be applied.
- The answers to this review must be on a separate sheet of lined paper. You will first see the objective of each question followed by the question itself. The question is what you are providing the answer to.

Objectives and Questions:

Chapter 1: Thinking Like a Scientist

- 1) The labs you do help you process and analyze information. They also teach you how to design labs. Be aware of how lab design works. Think critically about the errors that can occur.
 - a. Why are we do need multiple trials? Constants? Controls? Variables? Hypotheses?
 - b. Define accuracy and precision. It is possible to collect data that is both accurate and precise? Is it possible to collect data that is precise but not accurate?
 - c. What can you do when you run a lab to eliminate errors? Please make sure to refer to the equipment, the variables, and the materials used for the lab.
 - d. Can theories be modified? How has the atomic theory been modified over time?

Chapter 2: Atoms

- 1) Know how to identify a compound as an ionic or covalent compound. Covalent compounds are also referred to as molecular compounds.
 - a. What types of elements are in binary ionic compounds? Covalent Compounds (also called binary molecular compounds)?
 - b. Circle the following formulas that represent ionic compounds. Put a box around those that are covalent/molecular.

Pd_2S NaCl BaCl_2 O_2 FeSO_4 H_2O_2 Fe_2O_3 $(\text{NH}_4)_2\text{CO}_3$ PH_3
- 2) Know how to write the names and formulas for ionic compounds, covalent compounds, acids and bases.
 - a. Write the names for the following compounds:
 - i. SnO_2
 - ii. H_2S
 - v. KOH
 - vi. $\text{Mg}(\text{CN})_2$
 - ii. $\text{Fe}_3(\text{PO}_4)_2$
 - iv. HNO_3
 - vi. CO
 - viii. HNO_2
 - b. Write the formulas for the following compounds:
 - i. Silicon dioxide
 - ii. Phosphoric acid
 - iii. Lithium sulfite
 - iv. Tricarbon octahydride
 - v. potassium fluoride
 - vi. sodium hypochlorite
 - vii. Hydrocyanic acid
 - viii. Iron (II) oxide
 - c. Why do certain compounds require a roman numeral? What block of elements requires them?
- 3) Be able to predict the products of an acid and base reaction. Know what pH is and how to classify a solution as acidic, basic or neutral. Be familiar with common methods of pH detection.
 - a. Based on the following pH's, classify the solutions as acidic, basic, or neutral.
 - i. $\text{pH} = 2$
 - ii. $\text{pH} = 5.8$
 - iii. $\text{pH} = 7$
 - b. List two properties each for acids and bases. (Examples include taste, what ion they create, feel, and ability to react with a metal.)
 - c. What is more acidic, a solution with a pH of 2 or 5? Which is more basic, a pH of 8 or 14?

Utica Community Schools Chemistry Final Review

Daniela Niemeyer



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