

1. Prepping the crucibles and crucible lids. Observe, clean and dry two crucibles and their lids. As mentioned in the safety precautions, the crucibles will get hot when heated and feel changes color to marooned in the safety precautions, the crucibles with their lids on *you* use clay triangles. The lid should be opened first. Using tongs, place the crucibles with their lids on *you* use clay triangles. Then heat the crucibles with their lids slightly to allow water vapor to easily escape the crucible. Then heat the crucibles with their lids off for about four minutes. The point of this is to make sure the crucibles and their lids are completely dry. Turn the burner off and wait three minutes. **THE CRUCIBLES ARE STILL HOT!** Now, 20 minutes the tongs carry the crucibles to your bench. **THE CRUCIBLES ARE STILL HOT!** Now, 20 minutes while the crucibles and lids further cool. While the crucibles and lids are cooling proceed to step 2.
2. Obtain 1.0 grams of each hydrate and place each in a separate crucible, after the crucibles have cooled. $\text{CuSO}_4 \cdot x \cdot \text{H}_2\text{O}$ is the blue sample, while $\text{CoCl}_2 \cdot x \cdot \text{H}_2\text{O}$ is the purple sample. Put the crucible lids on top of the crucibles. Weigh each crucible, its lid and the hydrate, record these values in the data table at the top of page 3. In each sample x is an unknown integer.
3. Place each crucible with its lid on a clay support triangle and heat each for five minutes on low flame. After this slightly increase the flame and continue heating for ten minutes. After this time, turn off the burner and wait three minutes. **THE CRUCIBLES ARE STILL HOT!** Using tongs, carry the crucible burner and wait 8 minutes for the crucible to cool. Increasing the temperature slowly is to minimize the probability of spattering of the sample during dehydration. The lid is to prevent the loss of sample in case of some spattering. The cobalt chloride hydrate often does spatter if heated rapidly. Lastly, do not let the crucible get red hot since unwanted sample decomposition may occur.
4. Carry the crucible to the balance with tongs, the crucibles still will be slightly hot. Weigh each crucible lid and now anhydrous sample. Record these values.
5. Repeat steps three and four. Our goal here is to drive off any remaining water in the sample. If the difference in the masses between the two measurements is greater than 0.01 g, steps three and four must be repeated yet again. Record these values and all values in Table 1.
6. Keep a note of the colors of the samples before and after dehydration.

TABLE 1
DATA TABLE

Parameter	$\text{CuSO}_4 \cdot x \cdot \text{H}_2\text{O}$	$\text{CoCl}_2 \cdot x \cdot \text{H}_2\text{O}$
Mass of crucible and lid (g)	N/A	N/A
Mass of crucible, lid and sample (g)	N/A	N/A
Mass of sample (g)	1.000 g	1.000 g
Mass of crucible, lid and sample after first heating (g)	No Entry Here	No Entry Here
after second heating (g)	0.718 g	0.731 g
Mass of water lost (g)	0.282 g	0.269 g
Percent of water in sample by mass	28.2%	26.9%

Way To Go Indigo Lab Pre Lab Answers

**Edward Smedley, Hugh James
Rose, Henry John Rose**



Way To Go Indigo Lab Pre Lab Answers:

English Mechanic and Mirror of Science ,1909 Atlanta ,2004-05 Atlanta magazine s editorial mission is to engage our community through provocative writing authoritative reporting and superlative design that illuminate the people the issues the trends and the events that define our city The magazine informs challenges and entertains our readers each month while helping them make intelligent choices not only about what they do and where they go but what they think about matters of importance to the community and the region Atlanta magazine s editorial mission is to engage our community through provocative writing authoritative reporting and superlative design that illuminate the people the issues the trends and the events that define our city The magazine informs challenges and entertains our readers each month while helping them make intelligent choices not only about what they do and where they go but what they think about matters of importance to the community and the region The Carbon Age Eric Roston,2010-08-09 What do bubbles in a soft drink a bullet proof vest a plastic chair and our DNA have in common Carbon It is and forever has been the ubiquitous architect of life and civilization forming the chemical backbone of every living creature And yet when we hear the word today it is more often than not in a crisis situation carbon dioxide emissions are destroying the ozone layer and warming the planet the volatile Middle East explodes atop its stores of hydrocarbons carbohydrates threaten obesity and diabetics Carbon thus sustains us and threatens us in equal measure Eric Roston illuminates this essential element in all its forms cleverly recreating the intricate carbon cycle on the page by tracing its journey from the Big Bang to Earth and its extraordinary infiltration of this planet and in time influence on humankind and civilization Evoking its ubiquity more than 99% of all 31 million known substances contain carbon Roston chronicles the ways we have used it often to surprising and sometimes to catastrophic effect having sped up the carbon cycle in the last two centuries we are now attempting to wrestle Earth s geochemical cycle back from the brink Blending the latest science with original reporting Roston makes us aware as never before of the seminal impact carbon has and has had on our lives **Encyclopædia metropolitana; or, Universal dictionary of knowledge, ed. by E. Smedley, Hugh J. Rose and Henry J. Rose. [With] Plates** Encyclopaedia,1845 *American Druggists' Circular and Chemical Gazette* ,1892 The Northwestern Miller ,1919 **Encyclopaedia Metropolitana; Or, Universal Dictionary of Knowledge, on an Original Plan ... with ... Engravings: Mixed sciences** ,1845 Encyclopædia Metropolitana; Or, Universal Dictionary of Knowledge ... Edward Smedley,Hugh James Rose,Henry John Rose,1845 **The Louisiana Planter and Sugar Manufacturer** ,1897 **Encyclopaedia Metropolitana, Or, Universal Dictionary of Knowledge: Mixed sciences** Edward Smedley,Hugh James Rose,Henry John Rose,1845 **Outlook** Alfred Emanuel Smith,Francis Walton,1888

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