Name these Acids:

HIL hydroiodic acid mittric includ

H2501 H₁PO₄

sulfurous acid phosphoric acid

Write formulas for these acids:

hydrofluoric acid: _____ phosphorous acid: ______ phosphorous acid: ______

Name these bases and salts:

KOH potassium hydroxide

MqSO₄ magnesium sulfate

Calculate:

the pH of a 1.4 x 10⁻² M NaOH solution 1.

$$\begin{bmatrix} H^+ \end{bmatrix} \begin{bmatrix} OH^- \end{bmatrix} = I.0 \times I0^{-16} \qquad \begin{bmatrix} H^+ \end{bmatrix} \begin{bmatrix} I.4 \times I0^{-1} \end{bmatrix} = I.0 \times I0^{-16} \qquad \begin{bmatrix} H^+ \end{bmatrix} = 7.I \times I0^{-17} M$$

$$pH = -\log \begin{bmatrix} H^+ \end{bmatrix} = -\log \begin{bmatrix} 7.I \times I0^{-17} \end{bmatrix} = \boxed{pH = I2}$$

the [H'] of a solution with pH = 3.2

$$pH = -\log[H^*]$$
 $3.2 = -\log[H^*]$ $[H^*] = 6.3 \times 10^{-4} M$

the [OH] of a solution with a [H] of 9.3 x 10⁻⁴ M

Unit 12 Review Sheet Acids And Bases

J. G. R. Briggs

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