

Marking Scheme

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Page 1 Section A

1. (a) (i)



(1)

(ii) Same test

(1)

Calcium compounds give a brick-red flame.

(1)

(iii) (i) Isotopes are atoms of the same element with same proton number but different neutron numbers.

(1)

	Number of protons	Number of neutrons
$^{40}_{20}\text{Ca}$	20	20

(1)

(ii) (i) They have the same number of electrons in their outermost shells.

(1)

(ii) It has similar chemical properties to Ca does, then can replace some of the Ca required.

(1)

2. (a) (i) $2\text{AgCl} \rightarrow 2\text{Ag} + \text{Cl}_2$

(1)

(ii) The oxidation no. of Ag decreases and the oxidation no. of Cl increases.

(1)

(iii) No. of moles of $\text{AgCl} = 0.10 \times 1000/143.5 = 0.697$

No. of moles of $\text{Ag} = 2 \times$ No. of moles of AgCl

Mass of Ag that can be obtained = $100.9 \times$ No. of moles of Ag = $\frac{100.9 \times 0.697}{1000} = 0.703$

(1)

$= 0.703 \text{ g}$

(iv) (i) The black oxide changes to reddish brown metal

(1)

(ii) The metal obtained can conduct electricity.

(1)

(iii) $\text{CuO} + \text{H}_2 \rightarrow \text{Cu} + \text{H}_2\text{O}$

(1)

(iv) Hydrogen is explosive / flammable.

(1)

(v) No. The reactivity of Cu and Ag can only be compared using the same reaction.

(1)

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3. (a) (i) Sodium hydroxide is very caustic.

(1)

(ii) The sodium hydroxide is used.

(1)

(b) (i) Copper cannot displace H^+ (aq) from HCl (aq).

(1)

(ii) Add 2e^- from HCl (aq).

(1)

(c) (i) When water is added to concentrated H_2SO_4 , a lot of heat is produced. The heat can cause spilling out of the concentrated acid solution.

(1)

(ii) Add concentrated H_2SO_4 to water slowly and stir the mixture.

(1)

4. (a) (i) Chlorinate containing substances are harmful to marine life / water pollution.

(1)

(ii) (i) $\text{SO}_3^{2-} + \text{H}_2\text{O} \rightarrow \text{SO}_4^{2-} + 2\text{H}^+ + 2\text{e}^-$

(1)

(ii) $\text{Cr}_2\text{O}_7^{2-} + 14\text{H}^+ + 6\text{e}^- \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$

(1)

(iii) $\text{Cr}_2\text{O}_7^{2-} + 8\text{H}^+ + 6\text{SO}_3^{2-} \rightarrow 2\text{Cr}^{3+} + 6\text{SO}_4^{2-} + 7\text{H}_2\text{O}$

(1)

(iv) $\text{NaOH} + \text{Na}_2\text{CO}_3 + \text{HNO}_3 \rightarrow \text{NaNO}_3$

(1)

5. (a) Any TWO of the following points:

(1)

- Substances and series can be represented by a semi-empirical formula.

The general formula for pentane and octane is $\text{C}_n\text{H}_{2n+2}$. Adjacent members differ by one $-\text{CH}_2-$.

- There are gradual changes in physical properties among the members of a homologous series. The boiling point / melting point / density / refractive index is higher than that of pentane.

- Members of the same homologous series have similar chemical properties.

Both pentane and octane can undergo substitution reactions with Br_2/CCl_4 .

(b) Octane, C_8H_{18} , has a higher percentage of carbon by mass. Its chance to undergo incomplete combustion is greater than that of pentane.

(1)

(c) Any TWO of the following:

(1)

$\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$

$\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$

$\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$

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6. (a) (i)



(1)

(ii) It is a mixture of polymeric molecules of different chain lengths.

(1)

(iii) (i) condenser

(1)

(ii) Electric heating mantle / oil bath / sand bath should be used because benzene is flammable.

(1)

iii. Draw the skeleton of a branched polymer because styrene repeat is correct.

(1)

(iv) addition polymerisation

(1)

(v) (i) Electric heating can be prevented.

(1)

(ii) Plastic does not conduct easily.

(1)

(iii) Low density

(1)

7. Chemical Knowledge

(1)

Insert two of the metal strips into the beaker to form a chemical cell. Measure the voltage of the cell using the multimeter. Complete the electric circuit.

Control variables in the experiment such as:

- the size of the strips
- the separation / position between strips should be the same in each trial
- the temperature

The highest voltage can be obtained using a magnesium strip and a copper strip. (It is because among the three metals, Mg occupies the highest position in the e.m.f. and Cu the lowest position).

Adjust the distance / position between the two metal strips until the maximum voltage / current is obtained.

Apparatus / Concentration

(1)

Section B

8. (a)

$$\begin{aligned} \text{molar ratio} &= \frac{\text{Fe}}{\frac{100.1}{1000}} : \frac{\text{O}_2}{\frac{0.8}{32}} \\ &= 0.4375 : 0.025 \\ &= 17.5 : 1 \\ \text{Empirical formula of X is Fe}_{17.5}\text{O}_1 \end{aligned}$$

(b)

Let molar ratio of FeO to Fe_2O_3 be $x : y$

$$\frac{\text{molar ratio of Fe}}{\text{molar ratio of O}} = \frac{x + 2y}{x + 3y} = \frac{1}{4}$$

(1)

X is a mixture of FeO and Fe_2O_3 in a molar ratio of 2 : 1.

(1)

iii. X is not a mixture. In X, two-third of the total atoms is an oxidation number +2, while one-third is an oxidation number +3.

9. (a) (i) Hydrogen

(1)

(ii) manufacture of H_2SO_4 / manufacture of $-\text{CH}_2\text{OH}$ / manufacture of hydrochloric acid / fertilising of vegetable soils

(1)

(b) (i) Chlorine

(1)

(ii) At very low concentrations, Cl_2 will be formed + OH^- (aq) discharged.

(1)

Reaction of OH^- in HCl is higher than that of Cl^- / OH^- is a stronger reducing agent than Cl^- / OH^- loses electrons more readily than Cl^- .

(1)

(c) It is a reducing agent / It loses electrons

(1)

MnO_2 is an oxidising agent / MnO_2 gains electrons

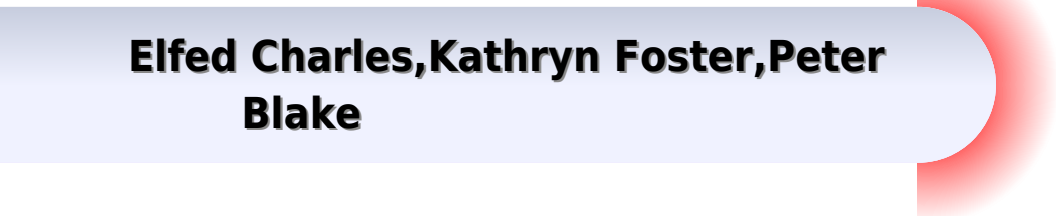
(1)

When the cell is connected in an external circuit, electrons will flow through the external circuit.

(1)

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**Elfed Charles, Kathryn Foster, Peter
Blake**



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