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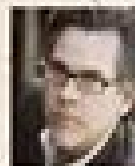
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WHAT BECOMES OF A CHARACTER CUT FROM A WRITER'S WORKING MANUSCRIPT?

On the eve of the attack on Pearl Harbor, Sam Sumida, a Japanese American academic, has been thrust into the role of amateur PI, investigating his wife's murder, which has been largely ignored by the LAPD. Grief-stricken by her loss, and disoriented by his ill-prepared change of occupation, Sam discovers that, inexplicably, not only has he become unrecognizable to his former acquaintances, but also all signs of his existence (including even the murder he's investigating) have been erased. Unaware that he is a discarded, fictional creation, he resumes his investigation in a world now characterized not only by his own sense of isolation but also by wartime fear.

To make matters worse, Sam finds himself on a collision course with a jingoistic and anti-Japanese PI—the revised, politically and commercially viable character for whom Sumida has been excised.

Behind it all is the ambitious, twenty-year-old Nisei author who has made the changes, despite his relocation to a Japanese internment camp. And looming above is his book editor in New York, who serves as both muse and manipulator to the young author—the woman with the blue pencil, a new kind of *femme fatale*.



GORDON McALPINE is the author of *Hammett Unwritten* and numerous other acclaimed novels, short stories, and nonfiction. He lives with his wife, Julie, in Southern California.



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WOMAN WITH A BLUE PENCIL

GORDON McALPINE

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"A brilliantly structured labyrinth of a novel... a totally unique work of 'mystery' fiction—one that Kafka, Borges, and Nabokov, as well as Dashiell Hammett, would have appreciated."
—JOYCE CAROL OATES

A Novel



GORDON McALPINE

Author of
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Woman With A Blue Pencil A Novel

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Taylor, Shirley Brooks, Francis Cowley
Burnand, Owen Seaman**

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web with a strong acid and strong base titration strong acids and strong bases refer to species that completely dissociate to form ions in solution by contrast weak acids and bases ionize only partially and the ionization reaction is reversible

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web aug 30 2022 acid base titrations are usually used to find the amount of a known acidic or basic substance through acid base reactions the analyte titrand is the solution with an unknown molarity the reagent titrant is the solution with a known molarity that will react with the analyte

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web conclusion the purpose of the experiment was to identify the unknown solid acid 17 by calculating its molar mass this was done by titration of the acid with sodium hydroxide in order to find the volume of sodium hydroxide required to reach the end point and that value was used in unit analysis to find the molar mass and identify the unknown solid acid

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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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web goals understand that parts of an acid base titration be able to determine the K_a or K_b from pH data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate K_{a1} and K_{a2} for a polyprotic acid by the end of this lab students should be able to

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web titration experiment in the neutralization of hydrochloric acid by sodium hydroxide the mole ratio of acid to base is 1:1
 $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$
 1 mole of HCl would be fully neutralized by one mole of NaOH

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web conclusion the identification of an unknown solid acid the purpose of this experiment was to determine the standardization of sodium hydroxide NaOH and oxalic acid $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$ and the molar mass of an unknown solid acid using the molar mass of the standardized sodium hydroxide solution

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web acid base reactions can be used to determine concentrations of an acid or a base using stoichiometric calculations and titrimetric analysis the volume of a solution required to react

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web jul 19 2023 introduction titration is an analytical quantitative technique used to determine the concentration of a solute a pH titration is used to determine the concentration of an acid or a base titrations play an important role in determining amount and purity in many manufacturing processes

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web specifically an acid base titration can be used to figure out the concentration of an acid or base whether an unknown acid or base is strong or weak pK_a of an unknown acid or pK_b of the unknown base key terms titration a process where a solution of known strength is added to a certain volume of a treated sample containing an

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web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being NaOH reaction with a 2nd substance $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{NaOH} \cdot \text{Na}_2\text{C}_2\text{O}_4 \cdot 4\text{H}_2\text{O}$ to differentiate an acid from a base in a solution a pH scale is used the pH scale ranges from 0 to 14 a pH of 7 is neutral in an acid base reaction

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web conclusion in this lab we used titration to explore the concepts of stoichiometry and equivalence points we found the concentration of an unknown substance by mixing 2M HCl with the NaOH of unknown concentration in order to experimentally ascertain the concentration of the NaOH

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web this point of titration is referred to as the equivalence point considering the 1:1 stoichiometry of this acid base reaction

NaOH aq C₆H₄COOH cook aq C₆H₄COONa cook aq H₂O l the point of equivalence is the point of titration when the number of moles of NaOH added is equal to the number of moles of KHP in the