

Chapter 5 Thermochemistry

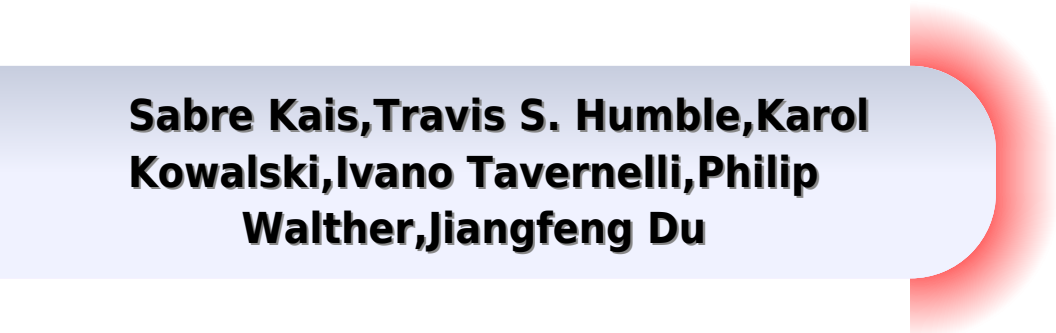
1. A kilogram of aluminum metal and a kilogram of water are each warmed to 75°C and placed in two identical insulated containers. One hour later, the two containers are opened, and the temperature of each substance is measured. The aluminum has cooled to 35°C . Assuming the identical insulated containers absorbed the same amount of energy, what would be the temperature of the water after one hour? Explain the differences in the final temperature of the water and aluminum.
2. Nitromethane (CH_3NO_2) burns in air (oxygen) to produce carbon dioxide, water, and nitrogen gas. If the $\Delta H^{\circ}_{\text{rxn}} = -1418 \text{ kJ}$, how much heat is produced by the complete reaction of 5.56 kg of Nitromethane?
3. A silver block initially at 58.5°C , is submerged into 100.0 g of water at 24.8°C , in an insulated container. The final temperature of the mixture upon reaching thermal equilibrium is 26.2°C . What is the mass of the silver block?
4. A 31.1 g of water of pure gold, initially at 69.3°C , is submerged into 64.2 g of water at 27.8°C in an insulated container. What is the final temperature of both substances at thermal equilibrium?
5. Zinc metal reacts with hydrochloric acid. When 0.105 g of Zn (s) reacts with 50.0 mL of 0.060 M HCl solution in a coffee cup calorimeter. The temperature of the solution raises from 22.5°C to 23.7°C . Find the $\Delta H^{\circ}_{\text{rxn}}$ for this reaction.
6. Instant cold packs used for sport injuries contain ammonium nitrate and water. Determine the enthalpy change $\Delta H^{\circ}_{\text{rxn}}$ for the dissolution of ammonium nitrate in water given the following. 1.25 g of ammonium nitrate is dissolved in 25.0 mL of water. The initial temperature is 25.8°C and the final temperature after dissolution is 21.9°C .
7. Calculate the $\Delta H^{\circ}_{\text{rxn}}$ for the reaction;

$$\text{CH}_4 (\text{g}) + 2 \text{O}_2 (\text{g}) \rightarrow \text{CO}_2 (\text{g}) + 2 \text{H}_2\text{O} (\text{g})$$

(1) $\text{C (s)} + 2 \text{H}_2 (\text{g}) \rightarrow \text{CH}_4 (\text{g})$	$\Delta H^{\circ}_{\text{rxn}} = ?$
(2) $2 \text{C (s)} + 2 \text{O}_2 (\text{g}) \rightarrow 2 \text{CO}_2 (\text{g})$	$\Delta H^{\circ}_{\text{rxn}} = -74.6 \text{ kJ}$
(3) $\text{H}_2\text{O (g)} \rightarrow \frac{1}{2} \text{O}_2 (\text{g}) + \text{H}_2 (\text{g})$	$\Delta H^{\circ}_{\text{rxn}} = -787 \text{ kJ}$
	$\Delta H^{\circ}_{\text{rxn}} = +241.8 \text{ kJ}$
8. Write the equation for the formation of each compound from its elements in their standard state ΔH°_f .
 - a. $\text{NH}_3 (\text{g})$
 - b. $\text{Fe}_2\text{O}_3 (\text{s})$
 - c. $\text{NaHCO}_3 (\text{s})$

Thermochemistry Guided Practice Problems

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