12.1 Ideal solutions

Solutions are among the most important kinds of systems in the chemical and biochemical sciences and technologies, spanning an enormous range of different phenomena and scales. The simplest model of a solution is the so-called ideal one.

An ideal solution is defined by the following properties.

- Components are non-interacting. In other words, there are no explicit energetic interactions between molecules of different species.
- (2) At constant temperature and pressure, an ideal entropy of mixing contributes to the total entropy due to the presence of different species.
- (3) Ideal solutions are usually excellent approximations to real ones in the near-pure limit, in which there is one major component (e.g., a solvent) and the concentrations of all other species (solutes) are very small.

To develop a thermodynamic model of ideal solutions, we first evaluate the entropy of mixing using a simple model. Figure 12.1 illustrates a lattice mixing model for a binary system. Though the lattice certainly does not represent molecular-structural features of actual liquids, the final behavior of the mixing entropy is surprisingly insensitive to its use. Prior to mixing the particles are completely separated into different regions of the system volume. One configuration exists for this state, and thus the entropy is zero. If the particles are then allowed to mix randomly, the number of arrangements of N_A A particles and N_B B particles in the $N=N_A+N_B$ sites is given by the combinations formula

$$\begin{split} &\Omega_{\text{mix}}(N_{\text{A}}, N_{\text{B}}) = \frac{(N_{\text{A}} + N_{\text{B}})!}{N_{\text{A}}! N_{\text{B}}!} \approx \frac{(N_{\text{A}} + N_{\text{B}})^{N_{\text{A}} + N_{\text{B}}}}{N_{\text{A}}^{N_{\text{A}}} N_{\text{B}}^{N_{\text{B}}}} \\ &= \left(\frac{N_{\text{A}}}{N_{\text{A}} + N_{\text{B}}}\right)^{-N_{\text{A}}} \left(\frac{N_{\text{B}}}{N_{\text{A}} + N_{\text{B}}}\right)^{-N_{\text{B}}} \end{split} \tag{12.1}$$

In the second line, we applied Stirling's approximation. By taking the logarithm to obtain the entropy, we find

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Oleinikovas, Aimee North, Haydn
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Fundamentals of Solidification 5th edition with Solutions Manual Wilfried Kurz, David J. Fisher, Michel Rappaz, 2023-08-23 Since the 4th 1998 edition there have been numerous crucial advances to the modelling and the basic understanding of solidification phenomena and with its linking to experimental results These topics have been incorporated into this 5th Fully Revised Edition as well as a new final chapter on microstructure selection which explains how to combine the concepts of the preceding chapters for modelling real microstructures in complex processes such as additive manufacturing This new 5th edition is of high interest to undergraduate and graduate levels and professionals With its numerous new topics also borne out by the new authorship students and teachers scientists and engineers will greatly benefit from this new book The topics are presented in the same praised manner as in previous editions readable at three levels an initial feel for the subject is obtained by consulting the figures and their detailed captions a deeper understanding of the underlying physics is found by working through the main text 15 appendices offer a detailed analysis of the various theories by providing detailed derivations of the relevant equations Particularly Novel the final chapter 8 on microstructure selection explains how to combine the concepts of the preceding chapters to model the real microstructures formed during complex processes such as additive manufacturing and the new detailed phase field appendix which opens the door to the accurate computer modelling of growth forms This edition goes with a companion Solutions Manual offering model solutions to 133 problems exercises **Biochemical Thermodynamics** Robert A. Alberty, 2006-03-31 Navigate the complexities of biochemical thermodynamics with Mathematica r Chemical reactions are studied under the constraints of constant temperature and constant pressure biochemical reactions are studied under the additional constraints of pH and perhaps pMg or free concentrations of other metal ions As more intensive variables are specified more thermodynamic properties of a system are defined and the equations that represent thermodynamic properties as a function of independent variables become more complicated This sequel to Robert Alberty's popular Thermodynamics of Biochemical Reactions describes how researchers will find Mathematica r a simple and elegant tool which makes it possible to perform complex calculations that would previously have been impractical Biochemical Thermodynamics Applications of Mathematica r provides a comprehensive and rigorous treatment of biochemical thermodynamics using Mathematica r to practically resolve thermodynamic issues Topics covered include Thermodynamics of the dissociation of weak acids Apparent equilibrium constants Biochemical reactions at specified temperatures and various pHs Uses of matrices in biochemical thermodynamics Oxidoreductase transferase hydrolase and lyase reactions Reactions at 298 15K Thermodynamics of the binding of ligands by proteins Calorimetry of biochemical reactions Because Mathematica r allows the intermingling of text and calculations this book has been written in Mathematica r and includes a CD ROM containing the entire book along with macros that help scientists and engineers solve their particular problems **Solutions Manual for an Introduction to Thermodynamics**

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