Study Guide: Thermal Energy

I. What is Thermal Energy?

- Definition: The total kinetic and potential energy of all the particles in a substance.
- Relationship to Particle Motion: The faster the particles move, the more thermal energy they have.
- Factors Affecting Thermal Energy:
 - Temperature
 - 6-September
 - State of matter (solid, liquid, gas)

III. Temperature

- Definition: A measure of the average kinetic energy of the particles in a substance.
- Units: Degrees Celsius (*C), degrees Fahrenheit (*F), Kelvin (kr)
- Relationship to Thermal Energy: Higher temperature means higher average kinetic energy, which generally means more thermal energy.

Mt. Heat

- Definition: The transfer of thermal energy from a warmer object to a cooler object.
- Direction of Heat Flows Heat always flows from hot to cold until thermal equilibrium is reached.
- Unit s: (oules (i), calories (cal).
 - Methods of Heat Transfer:
 - Conductions Heat transfer through direct contact.
 - · Example: Touching a hot stove.
 - Convection: Heat transfer through the movement of fluids (liquids or gases).
 - Example: Boiling water in a pot.
 - Radiation: Heat transfer through electromagnetic
 - Example: Feeling the heat from the sun.

IV. Thermal Properties of Matter

- Thermal Expansion and Contraction: Most substances expand when heated and contract when cooled.
 - Explanation: Particles move farther apart when heated and closer together when cooled.

Thermal Energy Study Guide

Raymond A. Serway, Steven Van Wyk

Thermal Energy Study Guide:

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