

Lab 1: *Microbiology: Media*

During this laboratory session, find the results of the reaction below below

Step 1: $\text{K}_2\text{Cr}_2\text{O}_7 + \text{H}_2\text{SO}_4 \rightarrow \text{H}_2\text{CrO}_4$

1. Balance the equation
2. If potassium dichromate is 295 grams initially and 15.00 grams of H_2SO_4 added to the flask, how much H_2CrO_4 is produced?
3. After finding the answer to #2, determine mass of H_2SO_4 that should be consumed (theoretical yield).
4. Convert it to grams initially following #2.
5. If this number is less than theoretical is it 100 grams of H_2SO_4 or less than 100 grams?

Unit 7 Stoichiometry Study Guide

Paul B. Kelter



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