

Titration and Neutralisation

Chemistry

Week 10

Name: _____

Date: _____

Acid-Base Titration

An acid-base **titration** is a technique by which the concentration of an acid or base is determined by reacting it with a known volume of a base or acid of known concentration. The **titrant** is the substance of known concentration that is added to the analyte. The **analyte** is the substance whose concentration is to be determined. The **equivalence point** is the point at which the acid and base have reacted in the mole ratio of their stoichiometric coefficients. The **end point** is the point at which the indicator changes color.

$$n(A) \cdot V(A) = n(B) \cdot V(B)$$

Equation

$n(A)$ is moles of acid

$V(A)$ is volume of acid

$n(B)$ is moles of base

$V(B)$ is volume of base

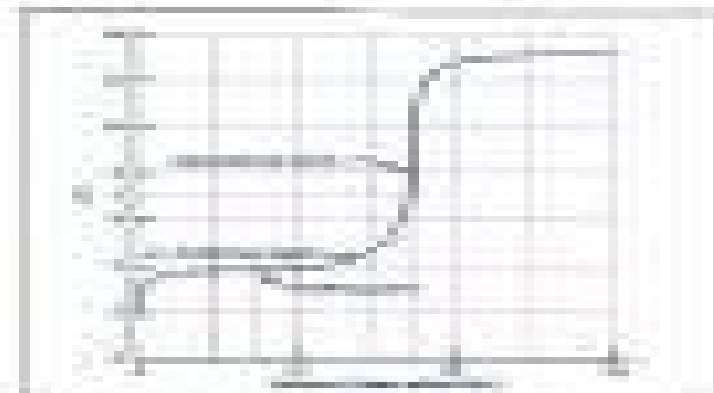
Acids and bases are substances that can donate or accept protons. Acids are substances that can donate protons, while bases are substances that can accept protons. The **equivalence point** is the point at which the acid and base have reacted in the mole ratio of their stoichiometric coefficients. The **end point** is the point at which the indicator changes color. The **equivalence point** is the point at which the acid and base have reacted in the mole ratio of their stoichiometric coefficients. The **end point** is the point at which the indicator changes color. The **equivalence point** is the point at which the acid and base have reacted in the mole ratio of their stoichiometric coefficients. The **end point** is the point at which the indicator changes color.

When an acid and a base are mixed, they react to form water and a salt. The **equivalence point** is the point at which the acid and base have reacted in the mole ratio of their stoichiometric coefficients. The **end point** is the point at which the indicator changes color. The **equivalence point** is the point at which the acid and base have reacted in the mole ratio of their stoichiometric coefficients. The **end point** is the point at which the indicator changes color.

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When the **equivalence point** is reached, the solution is neutral. The **equivalence point** is the point at which the acid and base have reacted in the mole ratio of their stoichiometric coefficients. The **end point** is the point at which the indicator changes color.

1. The **equivalence point** is the point at which the acid and base have reacted in the mole ratio of their stoichiometric coefficients. The **end point** is the point at which the indicator changes color.
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Equivalence point is the point at which the acid and base have reacted in the mole ratio of their stoichiometric coefficients.

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Lei Shi



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