Name:_____ Per:____

Unit 9 Outline- Thermodynamics, Reaction Rates, & Chemical Equilibrium Chapters 17 & 18

Essential Skills/ State Standards:

9-1. Describe how temperature and heat flow in terms of the motion of molecules (or atoms). (CSS: 7a)

9-2. Know that chemical processes & phase changes can either release energy (exothermic) or absorb energy (endothermic). (CSS: 7b. 7c)

9-3. Solve problems involving heat flow and temperature changes, using known values of specific heat (q= m · c · Δ T) & latent heat of phase change (q= m Δ H). (CSS) 7.69

9-4. Know the rate of reaction is the decrease in concentration of reactants or the increase in concentration of products with time & depends on such factors as: concentration, temperature, and pressure. (CSS: 8a: 8b)

9-5. Knew the definition and role of activation energy & catalyst in a chemical reaction. (CSS: 8d*, 8c)

9-6. Know how to use Le Chatelier's principle to predict the effect of changes in concentration, temperature, and pressure (CSS: 9a)

9-7. Know that equilibrium is established when the forward and reverse reaction rates are equal. (CSS: 9b)

Vecabulary.

- · Condensation
- Encezing
- Vaporization
- Melting.
- · Bailing
- Sublimation

- Deposition
- Heat
- Temperature
- Thermochemistry
- Exathermic
- Endothermic

- Calorimeter
- Joule
- Specific heat
- q=m·c·ΔT
- q= m △ H reperiorition
- q=m △H name

- Heat of reaction
- Le Chatlier's principle
- per wind opinion
- Equilibrium.
- Activistion Energy
- Catalyst

Book Assignments:

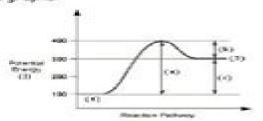
Topic	Read	Required Problems
I. Heat, Temperature & Specific Heat	Section 17-1 p. 511-524	Section 17-p. 524 #2, 5 & 6 Chapter 17 Review p. 548 #16 & 17
3. Chemical Equilibrium & Le Chatelier's Principle	Section 12-3 p. 372-375 Section 18-2 p. 562-568	Chapter 18 Review p. 586 #4, 6.6.7
4. Phase Changes Diagrams & Latent Heat of Phase Change	Section 12-3 p. 378-382	Section 12-3 p. 382 #4-8 Chapter 12 Review p. 389 #19, 28, 31 & 37
5. Collision Theory & Activation Energy	Section 17-3 p. 532-534	Section 17-3 p. 537 #2, 3 & 5
6. Reaction Rates & Catalysts	Section 17-4 p. 538-544	Section 17-4 p. 545 #2-4

Study Guide (Use notes, worksheets, and labs in addition to this study guide)

- 1. According to the collision theory, what 3 things must happen in order for two molecules to react?
- 2. What is the importance of the activation energy in a chemical reaction?
- Using the collision theory, explain if the following would increase/ decrease the rate of reaction:
 HCl(aq) + Mg(s) → MgCl₂(aq) + H₂(g)
 - a. Increase temp:
- b. Increase HCI
- c. Inchease pressure
- d. Remove MgCl₂
- e. Add He.

- 4. Identify the following terms with their correct number/letter from the graphs:
 - * Reactants-

- * Products-
- Activation energy (number/letter & amount in Joules)-
- * Is this reaction exothermic or endothermic?
- Explain how a catalyst affects the activation energy in a chemical reaction. (Also, show on the above graph what a catalyst might do).



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Margaret G. Kimble

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