

Chemistry/Chem Honors Unit 8 Test Study Guide

Part 1: Energy

- Recall the definition of energy.
- Recall the definitions of work and heat.
- Recall the symbols for work (w) and heat (q).
- Recall that the joule (J) is the SI unit for energy.
- Convert between joules, calories, and Calories (kcal).
- Define system and surroundings.
- Recall the definitions of endothermic and exothermic processes.
- Recall the first law of thermodynamics.
- Define heat capacity
- Define specific heat
- Convert between heat, specific heat, mass, and temperature change for a given process.
- Solve calorimetry problems similar to those in the ammonium nitrate lab and the HCl/NaOH lab.
- Recall that chemical reactions are accompanied by changes in enthalpy (ΔH).
- Recall that the sign of ΔH is negative for an exothermic reaction.
- Recall that the sign of ΔH is positive for an endothermic reaction.
- Determine the sign of ΔH based on a physical description of a given process.
- Recall that a balanced thermochemical equation relates ΔH for a reaction to the molar stoichiometric coefficients of all reactants & products.
- Convert between a balanced thermochemical equation, enthalpy, and moles or mass of a reactant or product.
- Describe Hess' law.
- Use Hess' law to determine the value of ΔH for a reaction by addition or subtraction of given reactions.
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