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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) + 2 \text{ C}_2\text{H}_2(\text{g}) + 5 \text{ O}_2(\text{g}) \rightarrow 2 \text{ H}_2\text{O}(\text{g}) + 2 \text{ C}_2\text{H}_6(\text{g})$ $\Delta H = 94.5 \text{ kJ}$
 $\text{H}_2\text{O}(\text{g}) + \text{H}_2(\text{g}) + \frac{1}{2} \text{ O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$ $\Delta H = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{g})$ $\Delta H = 283 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g})$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 2 \text{ H}_2\text{O}(\text{l})$ $\Delta H = 1411 \text{ kJ/mole}$
 $\text{C}_2\text{H}_6(\text{g}) + 7/2 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{l})$ $\Delta H = 1560 \text{ kJ/mole}$
 $\text{H}_2(\text{g}) + 1/2 \text{ O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$ $\Delta H = 285 \text{ kJ/mole}$

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web thus for reaction 9 2 in which 1 mole of h_2o and 1 mole of oh^- combine to form 2 moles of h_2o $\Delta h = -55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta h_{products} - \Delta h_{reactants}$ is positive if heat is absorbed that is if $\Delta h_{products}$

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web 3 12 hess s law in this experiment you will measure the amount of heat released in these three related exothermic reactions $1 \text{ NaOH (s)} \rightarrow \text{Na}^+ \text{ (aq)} + \text{OH}^- \text{ (aq)}$ $\Delta h = -44.5 \text{ kJ}$ $2 \text{ NaOH (s)} \rightarrow \text{H}^+ \text{ (aq)} + \text{Cl}^- \text{ (aq)}$ $\Delta h = -57.3 \text{ kJ}$ $2 \text{ H}_2\text{O (l)} \rightarrow \text{H}_2\text{ (g)} + \text{O}_2 \text{ (g)}$ $\Delta h = 571.6 \text{ kJ}$

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