

A romantic couple is shown in profile, smiling and looking at each other. They are sitting on a grassy field. The woman's hands are clasped together in the foreground, and the man's hands are resting on top of hers. The background is a soft, out-of-focus green field under bright light.

Always
AND
Forever

BAYOU DREAMS

FARRAH
ROCHON

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Xiaolong Qi



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calculated mass of the product expected $\times 100$ in this reaction a yield 80 is anticipated

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g Al $\times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 / 302.3 \text{ g Al}_2\text{O}_3 = 26.982 \text{ g Al}$ 4 mol Al 1 mol Al_2O_3 c determine the percent yield
yield actual $\times 100$ 260.0 g $\times 100$ 86.01 theoretical 302.3 g

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C is the limiting reagent 4.33 g of H₂ are left over 13

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CH₃OH can be produced by the following reaction $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$ calculate the theoretical yield of CH₃OH if 68.5 g of CO is
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 $80 \text{ g I}_2\text{O}_5 \times \frac{1 \text{ mol I}_2\text{O}_5}{1 \text{ mol I}_2} \times \frac{1 \text{ mol I}_2}{1 \text{ mol I}_2\text{O}_5} \times \frac{253.8 \text{ g I}_2}{1 \text{ mol I}_2} = 333.8 \text{ g I}_2$
 $28 \text{ g CO} \times \frac{1 \text{ mol CO}}{1 \text{ mol CO}} \times \frac{1 \text{ mol I}_2}{2 \text{ mol CO}} \times \frac{253.8 \text{ g I}_2}{1 \text{ mol I}_2} = 35.5 \text{ g I}_2$

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