

# Properties of water Study Guide with Summary questions

## Properties of Water



### Universal Solvent

Water is often referred to as the "universal solvent" due to its ability of solubility. This property is crucial for life on Earth and plays a key role in the biology of organisms.

The key to water's effectiveness as a solvent lies in its polarity, the positive charge on the hydrogen atoms and a partial negative charge on the oxygen atom. This polarity allows water to interact with and dissolve many different substances, the positive and negative charges of the water molecules effectively surrounding these dissolved particles. Because of its ability to dissolve substances, water serves as a medium for many biological processes and in the environment, its biological systems, its biochemical systems, facilitating the transport of nutrients, gases, and waste products.

Moreover, water's role as a solvent is crucial for life. It is essential for many biological processes, including those that involve the transport and the cycling of nutrients in ecosystems.

Water's versatility as a solvent also has practical applications in various fields, from environmental science to agriculture. Understanding its properties is essential for addressing environmental issues and for developing sustainable practices.

Moreover, it's important to note that water's ability to dissolve substances is not absolute. Some substances, such as oils and fats, are not soluble in water. Nevertheless, its universal solvent properties make water a critical factor shaping the Earth's environment and ecosystems.

Source: Adapted from "Properties of Water"

## Properties of Water

### Polarity

When we say that water is polar, we are referring to its molecular structure. Water molecules are composed of one oxygen atom and two hydrogen atoms. Oxygen is more electronegative than hydrogen, meaning it attracts electrons more strongly. This causes the shared electrons in the bond between oxygen and hydrogen to be pulled closer to the oxygen atom, thus creating a partial negative charge (delta minus) on the oxygen and a partial positive charge (delta plus) on the hydrogen atoms, giving water its polar nature.

This uneven distribution of charge makes water a polar molecule, where one end has a slightly negative charge and the other end has a slightly positive charge. You can think of it like a bar magnet, with its positive end being on the oxygen side and the negative end being on the hydrogen side. This polarity is responsible for many of water's unique properties, including its ability to act as a universal solvent and its high surface tension.

Water's polarity is crucial for many biological processes. It enables water molecules to form hydrogen bonds with each other and with other polar molecules or ions. Although these bonds are weak, they are strong enough to give water its unique characteristics, such as high surface tension, cohesion, and adhesion.

Through its polarity, water acts as a medium for many biological processes. Its partial charges on the oxygen and hydrogen atoms allow it to surround and interact with various ions and molecules, effectively separating them and allowing them to move freely. This ability makes water the "universal solvent" critical to all the major biological processes that sustain life on Earth.

In summary, the polarity of water results from the uneven distribution of electrons within the molecule, creating regions of partial positive and negative charge. This polarity is responsible for water's unique properties, such as its high surface tension and its ability to act as a universal solvent.



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Application of surface geophysics to ground-water investigations Geological Survey (U.S.), 1990      **Review of the Bureau of Reclamation's Corrosion Prevention Standards for Ductile Iron Pipe** National Research Council, Division on Engineering and Physical Sciences, National Materials Advisory Board, Committee on the Review of the Bureau of Reclamation's Corrosion Prevention Standards for Ductile Iron Pipe, 2009-11-19 Ductile iron pipe DIP was introduced about 50 years ago as a more economical and better performing product for water transmission and distribution As with iron or steel pipes DIP is subject to corrosion the rate of which depends on the environment in which the pipe is placed Corrosion mitigation protocols are employed to slow the corrosion process to an acceptable rate for the application When to use corrosion mitigation systems and which system depends on the corrosivity of the soils in which the pipeline is buried The Bureau of Reclamation s specification for DIP in highly corrosive soil has been contested by some as an overly stringent requirement necessitating the pipe to be modified from its as manufactured state and thereby adding unnecessary cost to a pipeline system This book evaluates the specifications in question and presents findings and recommendations Specifically the authoring committee answers the following questions Does polyethylene encasement with cathodic protection work on

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