

AP Chemistry Unit 4: Chemical Reactions
Worksheet 1

- Calculate the mass percent of oxygen in each of the following compounds:
(a) CuCO_3 (b) $\text{Fe}(\text{HCO}_3)_3$
- Calculate the mass of oxygen in each of the following samples:
(a) 2.45 g of $\text{C}_4\text{H}_8\text{O}_3$ (b) 7.10 g of $\text{Fe}(\text{HCO}_3)_3$
- A certain compound has a formula of Cu_xO_3 , where x is the subscript on the copper ion. The mass percent of the oxygen in a 2.0 g sample is found to be 11.1%.
(a) How many moles of oxygen are present in the sample?
(b) What is the molar mass of the compound?
(c) What is the subscript on the Cu?
$$6 \text{ Li}(s) + \text{N}_2(g) \rightarrow 2 \text{ Li}_3\text{N}(s)$$
- If 1.50 g of lithium reacts full on above the reaction:
(a) How many grams of N_2 also react?
(b) How many grams of Li_3N are created?
$$\text{MgCO}_3(s) + 2 \text{ HCl}(aq) \rightarrow \text{MgCl}_2(aq) + \text{CO}_2(g) + \text{H}_2\text{O}(l)$$
- When the above reaction goes to completion, 245 mL of CO_2 is produced at STP.
(a) What mass of CO_2 is present?
(b) What mass of MgCO_3 reacted?
(c) How many water molecules were produced?
$$2 \text{ NaNH}_2(s) \rightarrow 2 \text{ Na}(s) + \text{N}_2(g)$$
- 3.78×10^{23} molecules of N_2 are created when some NaNH_2 decomposes.
(a) How many moles of NaNH_2 decomposed?
(b) How many moles of Na are created?

Unit 4 Section 1 Quiz A Chemistry

Marcel A. Müller

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Table of Contents Unit 4 Section 1 Quiz A Chemistry

1. Understanding the eBook Unit 4 Section 1 Quiz A Chemistry
 - The Rise of Digital Reading Unit 4 Section 1 Quiz A Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Unit 4 Section 1 Quiz A Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Unit 4 Section 1 Quiz A Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Unit 4 Section 1 Quiz A Chemistry
 - Personalized Recommendations
 - Unit 4 Section 1 Quiz A Chemistry User Reviews and Ratings
 - Unit 4 Section 1 Quiz A Chemistry and Bestseller Lists
5. Accessing Unit 4 Section 1 Quiz A Chemistry Free and Paid eBooks

- Unit 4 Section 1 Quiz A Chemistry Public Domain eBooks
 - Unit 4 Section 1 Quiz A Chemistry eBook Subscription Services
 - Unit 4 Section 1 Quiz A Chemistry Budget-Friendly Options
6. Navigating Unit 4 Section 1 Quiz A Chemistry eBook Formats
- ePub, PDF, MOBI, and More
 - Unit 4 Section 1 Quiz A Chemistry Compatibility with Devices
 - Unit 4 Section 1 Quiz A Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Unit 4 Section 1 Quiz A Chemistry
 - Highlighting and Note-Taking Unit 4 Section 1 Quiz A Chemistry
 - Interactive Elements Unit 4 Section 1 Quiz A Chemistry
8. Staying Engaged with Unit 4 Section 1 Quiz A Chemistry
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Unit 4 Section 1 Quiz A Chemistry
9. Balancing eBooks and Physical Books Unit 4 Section 1 Quiz A Chemistry
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Unit 4 Section 1 Quiz A Chemistry
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Unit 4 Section 1 Quiz A Chemistry
- Setting Reading Goals Unit 4 Section 1 Quiz A Chemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Unit 4 Section 1 Quiz A Chemistry
- Fact-Checking eBook Content of Unit 4 Section 1 Quiz A Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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