

- 1.0×10^{-14} is the equilibrium constant for the autoionization of water (K_w)
- Equilibrium occurs in all solutions (acid, base, neutral)
 - All aqueous solutions have a detectable concentration of both hydronium and hydroxide ions
 - The product of all of these ion concentrations is always K_w
 - $K_w = \text{H}_3\text{O}^{+*} \text{OH}^{-} = 1.0 \times 10^{-14}$
- If concentration of hydroxide ion is known, to determine the pOH:
 - $\text{pOH} = 14 - \text{pH}$
- If pOH is known, to determine pH:
 - $\text{pH} = 14 - \text{pOH}$
- If the hydronium concentration is known, to determine the hydroxide ion concentration :
 - $(\text{OH}^{-}) = 1.0 \times 10^{-14} / (\text{H}_3\text{O}^{+})$
- Calculate pH from the hydronium concentration:
 - $\text{pH} = \log (\text{H}_3\text{O}^{+})$
- If pH is known, to determine the hydronium concentration:
 - $(\text{H}_3\text{O}^{+}) = 10^{-\text{pH}}$



Buffer Solutions

- A weak acid or base mixed with its conjugate
- It resists a change in pH when H^{+} or OH^{-} is added
- when H^{+} or OH^{-} is added to the buffer:
 - Determine which buffer component reacts
 - Work out the stoichiometry to find the new concentrations of buffer components
 - Use equilibrium constant to determine the new pH
- Ways to make buffers
 - Weak acid + Conjugate Base
 - Weak Base + Conjugate Acid
 - Excess Weak Acid + Strong Base
 - Excess Weak Base + Strong Acid

Unit 8 Chemistry Study Guide Answers

RD Boyd



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