

Name these Acids:

HI hydroiodic acid H_2SO_3 sulfurous acid HNO_3 nitric acid H_3PO_4 phosphoric acid

Write formulas for these acids:

hydrofluoric acid: HFphosphorous acid: H_3PO_3

Name these bases and salts:

KOH potassium hydroxide MgSO_4 magnesium sulfate

Calculate:

1. the pH of a
- 1.4×10^{-2}
- M NaOH solution

$$[\text{H}^+][\text{OH}^-] = 1.0 \times 10^{-14} \quad [\text{H}^+][1.4 \times 10^{-2}] = 1.0 \times 10^{-14} \quad [\text{H}^+] = 7.1 \times 10^{-13} \text{ M}$$

$$\text{pH} = -\log[\text{H}^+] = -\log[7.1 \times 10^{-13}] = \boxed{\text{pH} = 12}$$

2. the
- $[\text{H}^+]$
- of a solution with pH = 3.2

$$\text{pH} = -\log[\text{H}^+] \quad 3.2 = -\log[\text{H}^+] \quad \boxed{[\text{H}^+] = 6.3 \times 10^{-4} \text{ M}}$$

3. the
- $[\text{OH}^-]$
- of a solution with a
- $[\text{H}^+]$
- of
- 9.3×10^{-4}
- M

$$[\text{H}^+][\text{OH}^-] = 1.0 \times 10^{-14} \quad [9.3 \times 10^{-4}][\text{OH}^-] = 1.0 \times 10^{-14} \quad \boxed{[\text{OH}^-] = 1.1 \times 10^{-11}}$$

Unit 12 Review Sheet Acids And Bases

George Worthington

A red circular graphic with a gradient, appearing as a partial circle or a thick arc, located to the right of the author's name.

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