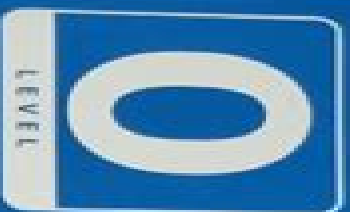


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TOPICAL EXAMINATION QUESTIONS

Chemistry



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ANSWERS (2011-2020)



TOPIC 1A: Experimental Design

Figure 1

Diagram shows a set-up for the experiment to determine the rate of reaction between hydrogen peroxide and potassium iodide. A gas syringe is used to measure the volume of oxygen gas produced. A known volume of hydrogen peroxide is added to a known volume of potassium iodide solution. The rate of reaction is determined by measuring the volume of oxygen gas produced at regular intervals.

Task 1: Design
Investigate the rate of reaction between hydrogen peroxide and potassium iodide.

The first student of the group found that the reaction was slow. A 10% of the reaction rate was found. The volume of oxygen gas produced was 10 cm³ after 10 minutes. A second student found that the reaction was faster. The volume of oxygen gas produced was 20 cm³ after 10 minutes. A third student found that the reaction was the fastest. The volume of oxygen gas produced was 30 cm³ after 10 minutes.

Task 2: Design
Investigate the rate of reaction between hydrogen peroxide and potassium iodide.

For a gas to be collected by the upward delivery method, it should be less dense than air. A gas to be collected by the downward delivery method, it should be denser than air.

Task 3: Design
Investigate the rate of reaction between hydrogen peroxide and potassium iodide.

1. A

Answer is an addition gas which can be used by using a gas syringe. The volume of gas produced is measured by the gas syringe. The rate of reaction is determined by measuring the volume of gas produced at regular intervals.

Task 4: Design
Investigate the rate of reaction between hydrogen peroxide and potassium iodide.

Using a known volume of hydrogen peroxide and a known volume of potassium iodide solution, the rate of reaction is determined by measuring the volume of oxygen gas produced at regular intervals. The rate of reaction is determined by measuring the volume of oxygen gas produced at regular intervals.

Task 5: Design
Investigate the rate of reaction between hydrogen peroxide and potassium iodide.

Use a known volume of hydrogen peroxide and a known volume of potassium iodide solution. The rate of reaction is determined by measuring the volume of oxygen gas produced at regular intervals.

Task 6: Design
Investigate the rate of reaction between hydrogen peroxide and potassium iodide.

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Benjamin Lillard



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Biomedical Engineering and Information Systems: Technologies, Tools and Applications Shukla, Anupam,Tiwari, Ritu,2010-07-31 Bridging the disciplines of engineering and medicine this book informs researchers clinicians and practitioners of the latest developments in diagnostic tools decision support systems and intelligent devices that impact and redefine research in and delivery of medical services Provided by publisher **Lumberman's Review** ,1920 **Answer Key for Physical Setting Chemistry** ,2008-09-01 Answer Key for past New York State Regents in Physical Setting Chemistry The American Produce Review ,1919 **Catalog of Copyright Entries. Third Series** Library of Congress. Copyright Office,1964 Includes Part 1 Number 1 Books and Pamphlets Including Serials and Contributions to Periodicals January June **Primary Education** ,1927 **New York Produce Review and American Creamery** ,1911 *Practical Druggist and Pharmaceutical Review of Reviews* Benjamin Lillard,1922 **American Druggist and Pharmaceutical Record** ,1901 **Grade Teacher** ,1928 **Energy Research Abstracts** ,1988 Practical Druggist and Pharmaceutical Review of Reviews ,1915

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