Name: ANSWER KEY

Chemistry 30 Unit 3: Chemical Equilibrium Assignment 2: 2-1 to 2-2

1. a) What is meant by a reversible reaction?

a reaction that can go in both forward and reverse directions.

- b) Are all chemical reactions reversible? No
- c) Are all reversible reactions always at equilibrium?
 No
- d) Does a reaction have to be reversible in order to reach equilibrium? Yes
- e) What, exactly, is equal at equilibrium? (define equilibrium)

When the rate of the forward reaction equals the rate of the reverse direction.

f) How is equilibrium different from a steady state system?

A steady state occurs in an open system, where there is continual addition of reactants and removal of products. An example would be an assembly line. An equilibrium occurs in a closed system, with no additional reactants being added or products being removed.

Write the equilibrium expression for each of the following reactions. Be sure to pay attention to physical states:

$$K_{eq} = \frac{[BeF_3]^2}{[Be_2][F_2]^5}$$

$$K_{eq} = \frac{[H_2O]^2[CI_2]^2}{[HCI]^4[O_2]}$$

$$K_{eq} = \frac{[Fe^{3+}]^5[Mn^{2+}]}{[Fe^{2+}]^5[MnO_4^-][H^+]^8}$$

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United States. National Bureau of Standards

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Salters Higher Chemistry, 1999 This work provides coverage of the content statements in the arrangements for Higher Chemistry organized by the three units in the course Energy Matters the World of Carbon and Chemical Reactions At the start of each unit students are given guidance on what they need to know and understand **Introduction to Hydrogeology** J.C. Nonner, Johannes Nonner, 2002-01-01 Providing an introduction to the crucially important topic of groundwater this text covers all major fields of hydrogeology and includes outlines of the occurrence of groundwater in various rock types the movement and storage of groundwater the formulation of groundwater balances the development of groundwater chemistry as well as the practical application of hydrogeology for groundwater development Following a unique systems approach to describe and connect its various elements the text also explores a large selection of examples of groundwater cases from various parts of the world In addition theoretical sections and examples are illustrated with a number of drawings photos and computer printouts Suitable for education in hydrogeology at postgraduate and graduate level the text is also a useful reference tool for professionals and decision makers involved in water or water related activities In the revised paperback edition of Introduction to Hydrogeology February 2006 suggestions of reviewers students and collaegues have been taken into account This means that more attention is paid to the processes in the unsaturated zone especially those relating to groundwater recharge Also in the revised edition the investigation methods are highlighted in the sections where the related theory is dealt with and they are not presented in the last chapter on groundwater management Chapter titles are re named and some definitions are adjusted The References and Bibliography section is also extended some figures are improved and the unevitable typing errors are corrected as well **Fundamental Experiments for College Chemistry** Harper Wine Frantz, Lloyd E. Malm, 1969 El-Hi Textbooks in Print ,1976 Excel Preliminary Chemistry C. M. Roebuck, 2003 Chemistry Education and Sustainability in the Global Age Mei-Hung Chiu, Hsiao-Lin Tuan, Hsin-Kai Wu, Jing-Wen Lin, Chin-Cheng Chou, 2012-12-05 This edited volume of papers from the twenty first International Conference on Chemical Education attests to our rapidly changing understanding of the chemistry itself as well as to the potentially enormous material changes in how it might be taught in the future Covering the full range of appropriate topics the book features work exploring themes as various as e learning and innovations in instruction and micro scale lab chemistry In sum the 29 articles published in these pages focus the reader's attention on ways to raise the quality of chemistry teaching and learning promoting the public understanding of chemistry deploying innovative technology in pedagogy practice and research and the value of chemistry as a tool for highlighting sustainability issues in the global community Thus the ambitious dual aim achieved in these pages is on the one hand to foster improvements in the leaching and communication of chemistry whether to students or the public and secondly to promote advances in our broader understanding of the subject that will have positive knock on effects on the world's citizens and environment In doing so the

book addresses as did the conference the neglect suffered in the chemistry classroom by issues connected to globalization even as it outlines ways to bring the subject alive in the classroom through the use of innovative technologies **Chemical Thermodynamics of Americium** R.J. Silva,G. Bidoglio,P.B. Robouch,I. Puigdomenech,H. Wanner,M.H. Rand,2012-12-02 This is the second volume in a series of critical reviews of the chemical thermodynamic data of those elements of particular importance in the safety assessment modeling of high level radioactive waste storage and disposal facilities The objective of these reviews is to provide a set of reliable thermodynamic data that can be used to describe the behaviour of these elements under conditions relevant for radioactive waste disposal systems and the geochemical environments The present volume is a review of experimental data reported in the literature for americium On a few occasions where no data existed comparisons and estimates were made based on experimental data on analog lanthanide elements The basic philosophy was to develop a minimum set of solid phases and solution species of americium that would fit all experimental data being reviewed

General Thermodynamics Donald Olander,2007-11-26 Because classical thermodynamics evolved into many branches of science and engineering most undergraduate courses on the subject are taught from the perspective of each area of specialization General Thermodynamics combines elements from mechanical and chemical engineering chemistry including electrochemistry materials science and biology to present a unique and thorough treatment of thermodynamics that is broader in scope than other fundamental texts This book contains classroom tested materials designed to meet the academic requirements for students from a variety of scientific and engineering backgrounds in a single course The first half focuses on classical concepts of thermodynamics whereas the latter half explores field specific applications including a unique chapter on biothermodynamics The book s methodology is unified concise and multidisciplinary allowing students to understand how the principles of thermodynamics apply to all technical fields that touch upon this most fundamental of scientific theories It also offers a rigorous approach to the quantitative aspects of thermodynamics accompanied by clear explanations to help students transition smoothly from the physical concepts to their mathematical representations Each chapter contains numerous worked examples taken from different engineering applications illustrations and an extensive set of exercises to support the material A complete solutions manual is available to professors with qualifying course adoptions

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