

Stoichiometry Study Guide

Students should be able to...

- Explain [mole ratio](#)
- Convert moles, grams, atoms/molecules
- Calculate theoretical yields when given the limiting reactant (Convert: [moles A to moles B](#), [grams A to moles B](#), [moles A to grams B](#), [grams A to grams B](#))
- Explain [limiting reactant and excess](#).
- Calculate theoretical yields by finding the [limiting reactant](#)
- Calculate [percent yield](#).
- Calculate percent composition.
- Calculate theoretical yield, limiting reactant, excess, and percent yield for a double replacement reaction (laboratory)

Practice Problems

Find the molar mass of each compound.

1. BaI_2
2. $\text{AgC}_2\text{H}_3\text{O}_2$

Convert.

3. 0.4 moles of Sn to atoms
4. 32 g of K_2O to molecules
5. 3.7 g of H_2 to moles
6. 800 g of CaCl_2 to moles
7. 0.75 moles of H_2O to grams
8. 9.0 g of Cl_2 to atoms
9. 0.5 moles of KNO_3 to grams

Find the percent composition for each element in the compound.

10. Mn_2O_7
11. AuNO_3

Stoichiometry



If you started with 3.7 moles of magnesium oxide, how many moles of magnesium hydroxide would be produced?

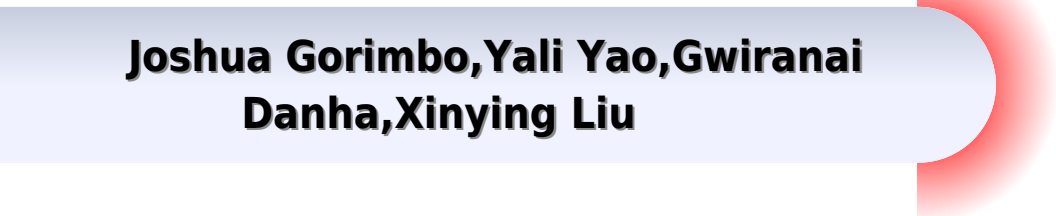


If 2.8 grams of sodium were used for the reaction, and an excess of iodine was present, how many moles of sodium iodide would be produced?



Unit 8 Stoichiometry Study Guide Answers

**Joshua Gorimbo, Yali Yao, Gwiranai
Danha, Xinying Liu**



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Robinson Chemistry Study Guide Robinson, 1992 **Chemistry, Study Guide** Bernice G. Segal, Peter S. Shenkin, 1989-02-14 This Second Edition of the first year chemistry text known for its clarity of exposition and its large number of illustrative worked problems contains a more rigorous treatment of electrochemistry chemical equilibrium and thermochemistry Worked examples now number over 300 and exercises over 1460 **Study Guide** Joan I. Senyk, Darrell D. Ebbing, 1984 *Study Guide for Chang's Chemistry* Kenneth W. Watkins, 1988 **General Chemistry Study Guide Sixth Edition** Darrell D. Ebbing, 1999 *Study Guide for General Chemistry and College Chemistry, Eighth Editions by Holtzclaw and Robinson* Norman E. Griswold, Henry Fuller Holtzclaw, 1988 **Study Guide** Darrell D. Ebbing, Susan M. Schelble, 1998

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included to help focus and inspire student interest frequent reference to common chemicals in commercial products numerous photographs of reactions in chapter practice exercises following worked examples **Ebook: Chemistry: The Molecular Nature of Matter and Change** Silberberg,2015-01-16 Ebook Chemistry The Molecular Nature of Matter and Change **Study Guide to Accompany Elements of Chemistry, General, Organic, and Biological -- Robert Boikess/Kenneth Breslauer/Edward Edelson** Robert A. Klein,1986 *10 in One Study Package for CBSE Chemistry Class 12 with Objective Questions & 3 Sample Papers 4th Edition* Disha Experts,2020-06-20 **Microbiology (Questions and Answers), 5e** Purshotam Kaushik & Kirti Kaushik,2022 Microbiology is an engaging textbook presenting balanced and comprehensive account of major areas of microbiology in the form of questions and answers This question answer approach to present complex topics and theories of microbiology regarding cellular and non cellular microorganisms microbial genetics and molecular biology in higher plants and animals makes the subject interesting and easily comprehensible for the students

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