Chemical Equilibrium

The Equilibrium Condition

- The majority of chemical reactions are reversible.
- Equilibrium is the state where the rate of the forward reaction is equal to the rate
 of the reverse reaction and results in no net observable change. At equilibrium,
 concentrations of all reactants and products remain constant at constant
 temperature.
- · Equilibrium is dynamic.

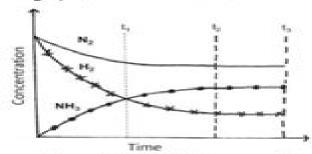
or -- indicates that the reaction is proceeding forward and in reverse.

- Once equilibrium is established, the rate of each direction is equal.
- This keeps the concentration of reactants constant
- This keeps the concentration of products constant
- This does NOT mean that concentration of products and concentration of reactants are the same or equal to each other.
- Any chemical reactions carried out in a closed vessel will reach equilibrium.
- Concept Check

$$3H_2(g) + 2N_2(g) \leftrightarrow 2NH_3(g)$$

H₂ gas and N₂ gas were placed in a rigid vessel and allowed to reach equilibrium in the presence of a catalyst according to the equation above. The diagram below shows how the concentrations of H₂, N₂, and NH₃ in this system changed over time. Based on the graph, which of these claims is most likely correct?





- a) The reaction reached equilibrium at t₁, because the concentrations of H₂ and NH₃ were the same.
- b) The reaction reached equilibrium at t₂, because the concentrations of N₂, H₂, and NH₃ stopped changing.
- c) The reaction reached equilibrium at t₃, because the reaction stopped proceeding.
- d) The reaction did not reach equilibrium, because not all of the N₂ and H₂ was consumed at time t₃.

Unit 4 Chemical Equilibrium Worksheet

Ray Sinnott, Gavin Towler

Unit 4 Chemical Equilibrium Worksheet:

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