

Stoichiometry Study Guide

Balance the following equation. Then use the balanced equation to answer the questions below.

11. If there are
a) How

b) How

c) How

12. If there are
a) How

b) How

c) How

13. If there are
a) How

b) How

c) How

Name _____ Period _____ Date _____

Stoichiometry Study Guide

All in the blanks below:

1 mole = _____ grams = _____ particles = _____ liters

Balance the following equation. Then use the balanced equation to answer the questions below.



11. If there are 2.8 moles of O_2 used in the reaction:

a) How many grams of NaCl would be needed to react?

b) How many grams of NaCl would be needed to react?

c) How many grams of NaCl would be needed to react?

d) How many grams of NaCl would be needed to react?

12. If the above chemical reaction is done as an experiment using 18 grams of each reactant, calculate how much NaClO₂ will be produced.

a) Given 18 grams of NaCl, how many moles of NaClO₂ will be produced?

b) Given 18 grams of O_2 , how many moles of NaClO₂ will be produced?

c) Which reactant is the limiting reactant? Which reactant is the excess reactant?

13. If the above chemical reaction is done as an experiment using 18 grams of each reactant, calculate the percent yield given that the actual yield of NaClO₂ is 35.8 grams.

a) What is the theoretical yield in grams of NaClO₂ given 18 grams of NaCl?

b) What is the percent yield given that the actual yield of NaClO₂ is 35.8 grams?

c) What is the theoretical yield in grams of NaClO₂ given 18 grams of O_2 ?

d) What is the percent yield given that the actual yield of NaClO₂ is 35.8 grams?

Unit 11 Stoichiometry Study Guide Answer Key

Leo J. Malone



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