

AP Chemistry Unit 4: Chemical Reactions
Worksheet 1

- Calculate the mass percent of oxygen in each of the following compounds:
(a) CuCO_3
(b) $\text{Fe}(\text{OH})_2$
- Calculate the mass of oxygen in each of the following samples:
(a) 2.45 g of $\text{C}_4\text{H}_8\text{O}_3$
(b) 7.10 g of $\text{Fe}(\text{OH})_2$
- A certain compound has a formula of Cu_xO , where x is the subscript on the copper ion. The mass percent of the oxygen in a 2.0 g sample is found to be 11.1%.
(a) How many moles of oxygen are present in the sample?
(b) What is the molar mass of the compound?
(c) What is the subscript on the Cu?
$$6 \text{ Li}(s) + \text{N}_2(g) \rightarrow 2 \text{ Li}_3\text{N}(s)$$
- If 1.50 g of lithium reacts fully via above the reaction:
(a) How many grams of N_2 also react?
(b) How many grams of Li_3N are created?
$$\text{MgCO}_3(s) + 2 \text{ HCl}(aq) \rightarrow \text{MgCl}_2(aq) + \text{CO}_2(g) + \text{H}_2\text{O}(l)$$
- When the above reaction goes to completion, 245 mL of CO_2 is produced at STP.
(a) What mass of CO_2 is present?
(b) What mass of MgCO_3 reacted?
(c) How many water molecules were produced?
$$2 \text{ NaN}_3(s) \rightarrow 6 \text{ Na}(s) + \text{N}_2(g)$$
- 3.78×10^{23} molecules of N_2 are created when some NaN_3 decomposes.
(a) How many moles of NaN_3 decomposed?
(b) How many moles of Na are created?

Unit 4 Section 1 Quiz A Chemistry

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