



Solubility Rules

Rule #		SOLUBILITY RULES
1	Cations	ALL Group 1 alkali metal (Li^+ , Na^+ , K^+ , Rb^+ , Cs^+ , Fr^+) and Ammonium (NH_4^+) compounds are <i>soluble</i> in water. Most, but NOT all, Group 2 alkali earth metal compounds are <i>soluble</i> .
2	Halides	Most Halides are <i>soluble</i> . Chlorides (Cl^-), Bromides (Br^-), and Iodides (I^-) are <i>soluble</i> EXCEPT Silver (Ag^+), Mercury (Hg_2^{2+}), and Lead (Pb^{2+})
3	Soluble Anions	Nitrates (NO_3^-), Acetates ($\text{C}_2\text{H}_3\text{O}_2^-$), Chlorate (ClO_3^-), and Perchlorates (ClO_4^-) are <i>soluble</i>
4	Sulfates	Most Sulfates (SO_4^{2-}) are <i>soluble</i> EXCEPT compounds of Barium (Ba^{2+}), Strontium (Sr^{2+}), Calcium (Ca^{2+}), Lead (Pb^{2+})
5	Insoluble Anions	Most Phosphates (PO_4^{3-}), Carbonates (CO_3^{2-}), Sulfides (S^{2-}), & Sulfites (SO_3^{2-}) are <i>insoluble</i> EXCEPT for compounds of Group 1 alkali metals and Ammonium (NH_4^+)
6	Hydroxides	Most hydroxides (OH^-) are <i>insoluble</i> EXCEPT compounds of Group 1 alkali metals, most Group 2 alkali earth metals, and Ammonium (NH_4^+)

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DP Hallahan



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