

6.9) It is often useful to write the equations for the reactions described:



The reaction for which the heat of reaction is requested is the formation of methane from its constituent elements.



The ΔH for this reaction can be calculated by summing the ΔH 's for a series of reactions which together add up to the same overall reaction (Hess's Law of Heat Summation).



These equations add up to the desired equation, so the ΔH values are added:

$$\Delta H^\circ(\text{CH}_4) = +890. - 394 - 572 \text{ kJ/mol} = -76 \text{ kJ/mol}$$

6.10) $\Delta H_{\text{rxn}} = -924 \text{ kJ}$

6.11) The heat of reaction is the difference between the heat of formation of the products and the heat of formation of the reactants in the balanced equation:

$$\Delta H^\circ = 12 \times (-384) + 6 \times (-242) - 2 \times (+49) = -8278 \text{ kJ}$$

There are two moles of benzene in the equation, so the heat released for one mole is 1/2 this amount, 3139 kJ/mol.

6.12) $\Delta H_{\text{rxn}} = -95.6 \text{ kJ}$

6.13) a. Heat lost = $50. \text{ g} \times 0.128 \text{ J/g}^\circ\text{C} \times 68^\circ\text{C} = 260 \text{ J}$

b. $q(\text{ice melting}) + q(\text{water warming}) = -q(\text{water cooling})$

$$5.0 \text{ g} \times 334 \text{ J/g} + 5.0 \text{ g} \times 4.184 \text{ J/g}^\circ\text{C} (T_f - 0^\circ\text{C}) = -100 \text{ g} \times 4.184 \text{ J/g}^\circ\text{C} (T_f - 70^\circ\text{C})$$

$$1670 \text{ J} + 20.92 T_f = -418.4 T_f + 29,288 \text{ J}$$

$$439.32 T_f = 27,618$$

$$T_f = 62.9 = 63^\circ\text{C}$$

6.14) a. $q(\text{water}) = -q(\text{coins})$

$$c \times m \times \Delta T = -c \times m \times \Delta T$$

$$4.184 \text{ J/g}^\circ\text{C} \times 200. \text{ g} \times 5.9^\circ\text{C} = -c \times 113.4 \text{ g} \times (-74.1^\circ\text{C})$$

$$\text{Solving for } c \text{ gives: } 0.567 \text{ J/g}^\circ\text{C}$$

b. No, the specific heat capacity of pure gold is $0.129 \text{ J/g}^\circ\text{C}$.

6.15) $\Delta H = 4.18 \text{ J/g}^\circ\text{C} \times 22.4 \text{ K} \times 70. \text{ g} = 6560 \text{ J}$

$$\text{For one mole of ammonium nitrate, } \Delta H = 6560 \times 60.05/20 = 26 \text{ kJ/mol}$$

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Iris Stovall, Donald J. DeCoste



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