

Versys ABS 2011



Specifications:

KLE650DBF

Engine Type	Inline Twin-Cylinder, 4-Stroke, Liquid-Cooled, DOHC 8-Valves
Displacement	649 cm ³
Bore & Stroke	83.0 x 60.0 mm
Compression Ratio	10.6:1
Fuel System	Fuel Injection, 38 mm x 2 with sub-throttles
Ignition	TCBI Digital Electronic Fuel Injection
Transmission	6-Speed, Cassette Style
Rake/Trail	25.0°/108 mm
Front Wheel Travel	150 mm
Rear Wheel Travel	145 mm
Front Tire Size	120/70-ZR17
Rear Tire Size	160/60-ZR17
Front Suspension	41mm Inverted Type Fork with Adjustable Preload and Rebound Damping RHD
Rear Suspension	Direct action offset rear shock absorber; Adjustable Spring Preload and Rebound Damping
Wheelbase	1,415 mm
Front Brake	Dual 300 mm semi-floating Petal Discs, Dual Opposed 4-pistons Callipers with ABS
Rear Brake	220 mm Single Petal Disc with Single Piston Calliper with ABS
Fuel Tank Capacity	19.0 litres
Ground Clearance	180 mm
Seat Height	815 mm
Gross Mass	209 kg (includes full fuel tank and all fluids at correct levels)
Colour	Metallic Spark Black with Super Flat Black

(Specifications subject to change without notice.)

Key Features:



649 cm³, Inline Twin-Cylinder, Liquid-Cooled



3-way Adjustable Windscreen



41 mm Inverted Forks, 300 mm Petal Discs, ABS Brakes



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web objective to apply stoichiometric principles to calculate the appropriate chemical equation for the reaction between iron metal and a solution of copper i sulfate introduction in the form of finely split powder copper metal is produced as a byproduct of the process as iron ions the ion meta dissolves

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web purpose to enhance the understanding of stoichiometry a reaction between iron and copper ii sulfate solution will be conducted this will help you to differentiate limiting and excess reactant in a chemical reaction finally the theoretical and percent yield of this reaction will be calculated

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web procedure 1 find the mass of a clean dry 100 ml beaker and record 2 measure out 8 00 grams of copper ii sulfate onto a

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web description great lab for use during stoichiometry unit in chemistry reinforces ideas of mole ratios and stoichiometry calculations by measuring the mass of iron that reacts and the mass of copper metal produced students can calculate the ratio of moles of reactant to moles of product this experimental mole ratio from lab can be compared

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web in this experiment we will use stoichiometric principles to deduce the appropriate equation for the reaction between metallic iron and a solution of copper ii sulfate this reaction produces metallic copper which is seen precipitating as a finely divided red powder

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web in this lab we will use experimental evidence and the principles of stoichiometry to determine which ion is formed in a single replacement reaction between iron and copper ii sulfate pre lab assignment in preparation for the lab activity we will investigate another single replacement reaction between copper and silver nitrate

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web through this experiment we sought to use stoichiometric principles which derive from stoichiometry the study of substance quantities in ratio or proportion in a chemical reaction in this lab experiment the principles help distinguish the equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction

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web this experiment is focused on a reaction between iron metal and a solution of copper ii sulfate we are able to figure out the chemical equation between the two by using stoichiometric principles the iron displaces the copper ii sulfate ions the copper ii is then reduced and gains electrons

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web stoichiometry can be defined as the relationship between the relative quantities of substances taking part in a reaction or forming a compound stoichiometric principles can be used to deduce the appropriate chemical equation for a reaction in our experiment we used solid iron in the form of powder was added to a solution of copper ii sulfate

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web under the finalize of the experiment it was found that iron was the limiting reactant and iron ii sulfate was a product of the reaction between 7 0002g away cooper ii sulfate and 2 0101g iron metal which produced 2 3037g of copper product with a percent yield of 100 72 introduction

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