

Stoichiometry Study Guide

Balance the following equation. Then use the balanced equation to answer the questions below.

11. If there are
a) 1 mole
b) 100 g

- a) 1 mole
b) 100 g

- a) 1 mole
b) 100 g

12. If there are
a) 1 mole
b) 100 g

- a) 1 mole
b) 100 g

- a) 1 mole
b) 100 g

13. If there are
a) 1 mole
b) 100 g

- a) 1 mole
b) 100 g

- a) 1 mole
b) 100 g

Name _____ Period _____ Date _____

Stoichiometry Study Guide

All in the blanks below:

1 mole = _____ grams = _____ particles = _____ liters

Balance the following equation. Then use the balanced equation to answer the questions below.



11. If there are 100 moles of O_2 used in the reaction:

a) How many grams of NaCl would be needed to react?

b) How many grams of NaClO_2 would be needed to react?

c) How many grams of NaCl would be needed to react?

d) How many grams of NaClO_2 would be needed to react?

12. If the above chemical reaction is done as an experiment using 10 grams of each reactant, calculate how much NaClO_2 will be produced.

a) Given 10 grams of NaCl , how many moles of NaClO_2 will be produced?

b) Given 10 grams of O_2 , how many moles of NaClO_2 will be produced?

c) Which reactant is the limiting reagent? Which reactant is the excess reagent?

13. If the above chemical reaction is done as an experiment using 10 grams of each reactant, calculate the percent yield given that the actual yield of NaClO_2 is 10.8 grams.

a) What is the theoretical yield in grams of NaClO_2 given 10 grams of NaCl ?

b) What is the percent yield given that the actual yield of NaClO_2 is 10.8 grams?

c) What is the theoretical yield in grams of NaClO_2 given 10 grams of O_2 ?

d) What is the percent yield given that the actual yield of NaClO_2 is 10.8 grams?

Unit 11 Stoichiometry Study Guide Answer Key

Puja Mehta



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