

Unit 6: Molecules & Compounds

Key Terms

Ionic Bonding

Electron dot diagram
ion
anion
cation
chemical bond
ionic bond
chemical formula
crystals

Covalent Bonding

Covalent bond
molecule
polar covalent bond

Naming Compounds & Writing Formulas

Polyatomic ion

Essential Questions

Electron Configurations: **electron dot diagram** – a model of an atom in which each dot represents a valence electron.

PERIODIC TABLE ELEMENTS 1–20							
H ·							He ·
Li ·	Be ·	B ·	C ·	N ·	O ·	F ·	Ne ·
Na ·	Mg ·	Al ·	Si ·	P ·	S ·	Cl ·	Ar ·
K ·	Ca ·						

• What is a chemical bond?

A **chemical bond** is the force that holds atoms and ions together as a unit.

• How is an ionic bond different from a covalent bond?

Ionic Bonds: some elements achieve stable electron configurations through the transfer of electrons between atoms.



Formation of ions – when an atom gains or losses an electron, the number of protons is no longer equal to the number of electrons. The charge is no longer neutral.

ion – an atom with a net positive or negative charge. Represented by a + or – sign.

Cl gains an electron 17 protons 18 electrons charge of 1- Cl^{1-} or Cl^-

anion – (an eye un) an ion with a negative charge

Naming: using part of the element name plus the suffix *-ide*.

Cl^- : *Chloride* ion.

Na loses an electron 11 protons 10 electrons charge of 1+ Na^{1+} or Na^+

cation – (kat eye un) an ion with a positive charge

Naming: just use the elements name.

Unit 6 Study Guide Covalent Bonding

Arihant Experts



Unit 6 Study Guide Covalent Bonding:

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Table of Contents Unit 6 Study Guide Covalent Bonding

1. Understanding the eBook Unit 6 Study Guide Covalent Bonding
 - The Rise of Digital Reading Unit 6 Study Guide Covalent Bonding
 - Advantages of eBooks Over Traditional Books
2. Identifying Unit 6 Study Guide Covalent Bonding
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Unit 6 Study Guide Covalent Bonding
 - User-Friendly Interface
4. Exploring eBook Recommendations from Unit 6 Study Guide Covalent Bonding
 - Personalized Recommendations
 - Unit 6 Study Guide Covalent Bonding User Reviews and Ratings
 - Unit 6 Study Guide Covalent Bonding and Bestseller Lists
5. Accessing Unit 6 Study Guide Covalent Bonding Free and Paid eBooks
 - Unit 6 Study Guide Covalent Bonding Public Domain eBooks
 - Unit 6 Study Guide Covalent Bonding eBook Subscription Services

- Unit 6 Study Guide Covalent Bonding Budget-Friendly Options
- 6. Navigating Unit 6 Study Guide Covalent Bonding eBook Formats
 - ePub, PDF, MOBI, and More
 - Unit 6 Study Guide Covalent Bonding Compatibility with Devices
 - Unit 6 Study Guide Covalent Bonding Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Unit 6 Study Guide Covalent Bonding
 - Highlighting and Note-Taking Unit 6 Study Guide Covalent Bonding
 - Interactive Elements Unit 6 Study Guide Covalent Bonding
- 8. Staying Engaged with Unit 6 Study Guide Covalent Bonding
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Unit 6 Study Guide Covalent Bonding
- 9. Balancing eBooks and Physical Books Unit 6 Study Guide Covalent Bonding
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Unit 6 Study Guide Covalent Bonding
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Unit 6 Study Guide Covalent Bonding
 - Setting Reading Goals Unit 6 Study Guide Covalent Bonding
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Unit 6 Study Guide Covalent Bonding
 - Fact-Checking eBook Content of Unit 6 Study Guide Covalent Bonding
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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