#### Name these Acids:

HIL hydroiodic acid mittric includ

H2501 H<sub>1</sub>PO<sub>4</sub>

sulfurous acid phosphoric acid

#### Write formulas for these acids:

hydrofluoric acid: \_\_\_\_\_ phosphorous acid: \_\_\_\_\_\_ phosphorous acid: \_\_\_\_\_\_

#### Name these bases and salts:

KOH potassium hydroxide

MqSO<sub>4</sub> magnesium sulfate

#### Calculate:

the pH of a 1.4 x 10<sup>-2</sup> M NaOH solution 1.

$$\begin{bmatrix} H^+ \end{bmatrix} \begin{bmatrix} OH^- \end{bmatrix} = I.0 \times I0^{-16} \qquad \begin{bmatrix} H^+ \end{bmatrix} \begin{bmatrix} I.4 \times I0^{-1} \end{bmatrix} = I.0 \times I0^{-16} \qquad \begin{bmatrix} H^+ \end{bmatrix} = 7.I \times I0^{-17} M$$
 
$$pH = -\log \begin{bmatrix} H^+ \end{bmatrix} = -\log \begin{bmatrix} 7.I \times I0^{-17} \end{bmatrix} = \boxed{pH = I2}$$

the [H'] of a solution with pH = 3.2

$$pH = -\log[H^*]$$
  $3.2 = -\log[H^*]$   $[H^*] = 6.3 \times 10^{-4} M$ 

the [OH] of a solution with a [H] of 9.3 x 10<sup>-4</sup> M

# **Unit 12 Review Sheet Acids And Bases**

**Annu Kaushik** 

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- 1. Understanding the eBook Unit 12 Review Sheet Acids And Bases
  - The Rise of Digital Reading Unit 12 Review Sheet Acids And Bases
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Unit 12 Review Sheet Acids And Bases
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Unit 12 Review Sheet Acids And Bases
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Unit 12 Review Sheet Acids And Bases
  - Personalized Recommendations
  - Unit 12 Review Sheet Acids And Bases User Reviews and Ratings
  - Unit 12 Review Sheet Acids And Bases and Bestseller Lists
- 5. Accessing Unit 12 Review Sheet Acids And Bases Free and Paid eBooks
  - Unit 12 Review Sheet Acids And Bases Public Domain eBooks
  - Unit 12 Review Sheet Acids And Bases eBook Subscription Services
  - Unit 12 Review Sheet Acids And Bases Budget-Friendly Options
- 6. Navigating Unit 12 Review Sheet Acids And Bases eBook Formats

- ePub, PDF, MOBI, and More
- Unit 12 Review Sheet Acids And Bases Compatibility with Devices
- Unit 12 Review Sheet Acids And Bases Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Unit 12 Review Sheet Acids And Bases
  - Highlighting and Note-Taking Unit 12 Review Sheet Acids And Bases
  - Interactive Elements Unit 12 Review Sheet Acids And Bases
- 8. Staying Engaged with Unit 12 Review Sheet Acids And Bases
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - o Following Authors and Publishers Unit 12 Review Sheet Acids And Bases
- 9. Balancing eBooks and Physical Books Unit 12 Review Sheet Acids And Bases
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Unit 12 Review Sheet Acids And Bases
- 10. Overcoming Reading Challenges
  - $\circ\,$  Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Unit 12 Review Sheet Acids And Bases
  - Setting Reading Goals Unit 12 Review Sheet Acids And Bases
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Unit 12 Review Sheet Acids And Bases
  - Fact-Checking eBook Content of Unit 12 Review Sheet Acids And Bases
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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