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 $2 \text{ CO}_2(g) + \text{H}_2\text{O}(g) \rightarrow \text{C}_2\text{H}_2(g) + \frac{5}{2} \text{O}_2(g)$ $\Delta H = 94.5 \text{ kJ}$
 $\text{H}_2\text{O}(g) + \frac{1}{2} \text{O}_2(g) \rightarrow \text{H}_2\text{O}(l)$ $\Delta H = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(g) + \frac{7}{2} \text{O}_2(g) \rightarrow 2 \text{ CO}_2(g) + 3 \text{ H}_2\text{O}(g)$ $\Delta H = 283 \text{ kJ}$
 $\text{C}_2\text{H}_6(g)$

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 $\text{C}_2\text{H}_4(g) + 3 \text{ O}_2(g) \rightarrow 2 \text{ CO}_2(g) + 2 \text{ H}_2\text{O}(l)$ $\Delta H = 1411 \text{ kJ/mole}$
 $\text{C}_2\text{H}_6(g) + \frac{7}{2} \text{ O}_2(g) \rightarrow 2 \text{ CO}_2(g) + 3 \text{ H}_2\text{O}(l)$ $\Delta H = 1560 \text{ kJ/mole}$
 $\text{H}_2(g) + \frac{1}{2} \text{ O}_2(g) \rightarrow \text{H}_2\text{O}(l)$ $\Delta H = 285 \text{ kJ/mole}$

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