\_22.5 L O<sub>2</sub>\_\_\_\_10. Calculate the number of liters of oxygen gas needed to produce 15.0 liters of Dinitrogen trioxide.

N <sub>2</sub>	+	302	-	2N-0
		+		+
		mol		mol

15.0 LN203	1 mol-N <sub>2</sub> O <sub>3</sub>	3 mgl-02	22.4 L O
	22.4 J. N2O3	2 mol N2O3	1 mp4-0 ≤

Write the correct answer for each of the following statements or questions in the blank provided. Use your book, if necessary, to find the terms that fit the definitions.

6.02 X 10 <sup>23</sup>	1.1	Avogadro's number?
-------------------------	-----	--------------------

Reactants	11.50	the starting materials in an equation.
Mass-mass		
Calculations	16.	determining the mass of products from the given mass of
	100 a C	tants

Multiple Choice

- \_B\_\_17. Which of these expressions is an <u>incorrect</u> interpretation of the balanced equation? 2S + 3O<sub>2</sub> → 2SO<sub>3</sub>
  - A. 2 atoms S + 3 molecules O₂ → 2 molecules S O₃
  - B.  $2qS + 3qO_2 \rightarrow 2qSO_3$
  - C. 2 mole S + 3 mole O₂ → 2 mole SO₃
  - D. none of the above
- \_\_\_\_\_\_\_18. In a chemical reaction, the mass of the products:
  - is less than the mass of the reactants.
  - Is greater than the mass of the reactants.
  - Is equal to the mass of the reactants.
  - D. Has no relationship to the mass of the reactants.

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