

BALANCING AND TYPES OF REACTIONS QUIZ

- Chemical equations must be balanced to satisfy _____.
 A) the law of definite proportions C) the law of conservation of mass
 B) the law of multiple proportions D) Avogadro's principle
- In every balanced chemical equation, each side of the equation has the same number of _____.
 A) atoms of each element B) molecules C) Moles D) coefficients
- A chemical equation is balanced when the:
 A) coefficients of the reactants equal the coefficients of the products.
 B) same number of each kind of atom appears in the reactants and in the products.
 C) products and reactants are the same chemicals.
 D) subscripts of the reactants equal the subscripts of the products.
- The complete balanced equation for the reaction between zinc hydroxide and acetic acid is
 A) $\text{ZnOH} + \text{CH}_3\text{COOH} \rightarrow \text{ZnCH}_3\text{COO} + \text{H}_2\text{O}$.
 B) $\text{Zn}(\text{OH})_2 + \text{CH}_3\text{COOH} \rightarrow \text{Zn} + 2\text{CO}_2 + 3\text{H}_2\text{O}$.
 C) $\text{Zn}(\text{OH})_2 + 2\text{CH}_3\text{COOH} \rightarrow \text{Zn}(\text{CH}_3\text{COO})_2 + 2\text{H}_2\text{O}$.
 D) $\text{Zn}(\text{OH})_2 + 2\text{CH}_3\text{COOH} \rightarrow \text{Zn}(\text{CH}_3\text{COO})_2 + \text{H}_2 + \text{O}_2$.
- Which equation is not balanced?
 A) $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ B) $4\text{H}_2 + 2\text{O}_2 \rightarrow 4\text{H}_2\text{O}$
 C) $\text{H}_2 + \text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}$ D) $2\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$
- When the following equation is balanced, the coefficients are _____.
 $\text{C}_3\text{H}_8 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
 A) 2, 3, 4, 4 B) 1, 4, 8, 9 C) 2, 12, 8, 9
 D) 4, 4, 32, 36 E) 2, 25, 16, 18
- When the following equation is balanced, the coefficient of water is _____.
 $\text{N}_2\text{O}_5 (\text{g}) + \text{H}_2\text{O} (\text{l}) \rightarrow \text{HNO}_3 (\text{aq})$
 A) 1 B) 2 C) 3 D) 4 E) 5
- A group whose atoms (M) undergo the reaction:
 $\text{MO}(\text{s}) + \text{H}_2\text{O}(\text{l}) \rightarrow \text{M}^{2+}(\text{aq}) + 2\text{OH}^{-}(\text{aq})$
 A) group 1 B) group 2 C) group 13 D) group 14 E) group 15
- State the type of reaction: $2\text{K} + 2\text{H}_2\text{O} \rightarrow 2\text{KOH} + \text{H}_2$
 A) synthesis B) decomposition C) single displacement
 D) double displacement E) combustion F) no reaction
- State the type of reaction: $4\text{Cr} + 3\text{O}_2 \rightarrow 2\text{Cr}_2\text{O}_3$
 A) synthesis B) decomposition C) single displacement
 D) double displacement E) combustion F) no reaction

Types Of Reactions Quiz Answers

Ben Greenstein



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