

Name: _____
Date: _____ Per. #: _____

Chemistry – Unit 8

Unit 8 Exam Study Guide

NOTE: *This is a 30 point homework assignment!*

Topics

Lesson 8.1 – Oxidation and Reduction (pgs. 631-635)

- Know all the ways to determine if an element is being oxidized or reduced.
- Be able to assign oxidation states to all elements in a chemical reaction.
- Be able to use oxidation states to determine if a reaction is redox or not.
- Be able to turn word equations into symbolic equations.
- Be able to add half reactions to create a balanced overall redox reaction.

Lesson 8.2 – Balancing Redox Reactions (pgs. 637-641)

- Be able to determine the number of electrons being gained and lost in an unbalanced redox reaction.
- Be able to balance a redox reaction using the half reaction method. Memorize the steps!
- Remember it is ok and even helpful to leave out spectator ions (ions that do not change oxidation state and therefore are not a part of the redox reaction).

Lesson 8.3 – Electrochemical Cells (pgs. 655-657, 663, 665)

- Know the parts of an electrochemical cell and be able to describe or even draw a picture of one.
- Know the difference between a voltaic (or galvanic) cell and an electrolytic cell.
- Know how to identify the anode and the cathode.
- Know how electrons would travel in the wire and how ions would migrate across a salt bridge in a cell.
- Know the meaning of the standard reduction potential (SRP) and be able to find one for any half-reaction.
- Know how to use SRPs to calculate the potential of an oxidation reaction and the overall cell potential (E°_{cell}).
- Know how to tell if a redox reaction would be spontaneous or not based on the position of the half-reactions on the SRP handout and the sign of E°_{cell} .
- Know how to write the cell notation for any electrochemical cell.

Lab 8.1 – Electrolytic Cells (pgs. 667-671)

- Know how an overpotential is necessary to drive a nonspontaneous reaction in an electrolytic cell.
- Know how to choose between competing possibilities for the oxidation and reduction half reactions to determine which ones would occur in an electrolytic cell.

Practice Problems

1. For each of the following, write "R" if the statement describes reduction or "O" if the statement describes oxidation.

_____ Occurs at the anode

_____ $\text{Cu}^{2+} + 2\text{e}^- \rightarrow \text{Cu}^0$

_____ Gains electrons

_____ Charge on element becomes more negative

_____ Electrons on the right side of arrow

_____ $2\text{F}^- \rightarrow \text{F}_2 + 2\text{e}^-$

_____ Loses electrons

_____ $\text{Cl}_2 + 2\text{e}^- \rightarrow 2\text{Cl}^-$

_____ Occurs at the cathode

_____ Charge on element becomes more positive

Unit 8 Study Guide Chemistry Answer Key

RJ Shavelson



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Organic Chemistry Study Guide Robert J. Ouellette, J. David Rawn, 2014-11-04 Organic Chemistry Study Guide Key Concepts Problems and Solutions features hundreds of problems from the companion book Organic Chemistry and includes solutions for every problem Key concept summaries reinforce critical material from the primary book and enhance mastery of this complex subject Organic chemistry is a constantly evolving field that has great relevance for all scientists not just chemists For chemical engineers understanding the properties of organic molecules and how reactions occur is critically important to understanding the processes in an industrial plant For biologists and health professionals it is essential because nearly all of biochemistry springs from organic chemistry Additionally all scientists can benefit from improved critical thinking and problem solving skills that are developed from the study of organic chemistry Organic chemistry like any skill is best learned by doing It is difficult to learn by rote memorization and true understanding comes only from concentrated reading and working as many problems as possible In fact problem sets are the best way to ensure that concepts are not only well understood but can also be applied to real world problems in the work place Helps readers learn to categorize analyze and solve organic chemistry problems at all levels of difficulty Hundreds of fully worked practice problems all with solutions Key concept summaries for every chapter reinforces core content from the companion book

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