

Unit 6: Molecules & Compounds

Key Terms

Ionic Bonding

Electron dot diagram
ion
anion
cation
chemical bond
ionic bond
chemical formula
crystals

Covalent Bonding

Covalent bond
molecule
polar covalent bond

Naming Compounds & Writing Formulas

Polyatomic ion

Essential Questions

Electron Configurations: **electron dot diagram** – a model of an atom in which each dot represents a valence electron.

PERIODIC TABLE ELEMENTS 1–20							
H ·							He ·
Li ·	Be ·	B ·	C ·	N ·	O ·	F ·	Ne ·
Na ·	Mg ·	Al ·	Si ·	P ·	S ·	Cl ·	Ar ·
K ·	Ca ·						

• What is a chemical bond?

A **chemical bond** is the force that holds atoms and ions together as a unit.

• How is an ionic bond different from a covalent bond?

Ionic Bonds: some elements achieve stable electron configurations through the transfer of electrons between atoms.



Formation of ions – when an atom gains or losses an electron, the number of protons is no longer equal to the number of electrons. The charge is no longer neutral.

ion – an atom with a net positive or negative charge. Represented by a + or – sign.

Cl gains an electron 17 protons 18 electrons charge of 1- Cl^{1-} or Cl^-

anion – (an eye un) an ion with a negative charge

Naming: using part of the element name plus the suffix *-ide*.

Cl^- : *Chloride* ion.

Na loses an electron 11 protons 10 electrons charge of 1+ Na^{1+} or Na^+

cation – (kat eye un) an ion with a positive charge

Naming: just use the elements name.

Unit 6 Study Guide Covalent Bonding

**Marcia L. Gillette, Ira Blei, Wendy
Gloffke, George Odian**



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Study Guide to Accompany Basics for Chemistry Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

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