

Synthesis of Alum, $KAl(SO_4)_2 \cdot 12H_2O$

The purpose of this experiment is to synthesize potassium aluminum sulfate, $KAl(SO_4)_2 \cdot 12H_2O$, commonly referred to as "alum".

Mass approximately 1.8 grams of Aluminum foil. Add 20 mL KOH.

Transfer this solution and pour back into beaker.

Slowly add 20 mL Sulfate Acid and filter again.

Allow crystals to form. Filter again with filtered glass product.

Quantitative

Mass of aluminum foil	0.999 g
Mass of filter paper (3 pieces)	3.015 g
Mass of fp. and alum	11.194 g
Mass of alum crystals	8.179 g

Qualitative

The solution bubbles when KOH is added. The precipitate formed is white and chalky. Alum Crystals are numerous, transparent, and come in a wide variety of shapes.

Draw a picture of the alum crystals here.



1. Write balanced net ionic equations for the following reactions which occur in this synthesis:

- a. aluminum reacts with KOH and water forming potassium ions, $[Al(OH)_4]^-$ and hydrogen gas



- b. hydrogen ions from the acid react with the tetrahydroxaluminate ions to precipitate aluminum hydroxide



- c. aluminum hydroxide reacts with additional hydrogen ions and water to form hydrated aluminum ions $[Al(H_2O)_6]^{3+}$



- d. alum forms from the potassium ions, hydrated aluminum ions, sulfate ions, and water



2. Calculate the theoretical yield of alum. What assumptions are being made in your determination of the theoretical yield?

$$(0.999 \text{ g Al} / 1) * (1 \text{ mol Al} / 26.982 \text{ Al}) * (1 \text{ mol Alum} / 1 \text{ mol Al}) * (474.338 \text{ g Alum} / 1 \text{ mol Alum}) = 17.6 \text{ g Alum}$$

We are assuming that the Aluminum foil is pure Aluminum and that all impurities were removed in the filtering process.

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Raymond A. Serway, John W. Jewett



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