

3.3 Thermodynamics: Study Guide

Use this [outline](#) to review the key vocabulary for this unit.

1. Match the vocabulary words on the left with the appropriate definitions and units on the right.

a. Thermal Energy	<u>A</u> Kinetic energy in transit from one object to another due to temperature difference
b. Temperature	<u>B</u> Degree Celsius
	<u>C</u> Joules
	<u>D</u> The average kinetic energy of particles in an object.

2. In the table below, write in the appropriate formulas.

Energy Transfer	Rate of Energy Transfer	Thermal Conductivity	Coefficient of Heat Conductivity
$Q = mc \cdot \Delta T$	$P = \frac{c \Delta T}{L}$ or $P = \frac{Q}{t}$ <u>Thermal conductance</u> $P = \sigma A c (\Delta T)^2$	$K = \frac{PL}{A \Delta T}$	$U = \frac{P}{A \cdot \Delta T}$

3. For each example, write the correct number for the law of thermodynamics it describes (0, 1, 2).
- 2 The law of entropy
 - 1 Thermal energy can only be increased in a system by adding heat from an external source or performing work within the system
 - 0 If two systems are separately found to be in thermal equilibrium with a third, then the two systems are in thermal equilibrium with each other.
 - 2 Thermal energy flows only from warmer to cooler
 - 1 The law of conservation of energy applied to thermal systems
4. How much thermal energy would it take to raise the temperature of 1.1 (1 kg) of water by 10°C? The specific heat capacity of water is $4184 \frac{J}{kg \cdot ^\circ C}$

$$Q = mc \Delta T$$

$$= 1.1 kg \cdot 4184 \frac{J}{kg \cdot ^\circ C} \cdot 10^\circ C$$

$$Q = 41,840 J$$

$$Q = \underline{41,840 J}$$

5. A materials scientist is trying to determine the specific heat capacity of an experimental material. She applies 30 kJ of thermal energy to a 5 kg sample of the material and records a 24°C temperature increase in the sample. What is the specific heat capacity of the material?

$$Q = mc \Delta T$$

$$c = \frac{Q}{m \Delta T}$$

$$c = 2500 \frac{J}{kg \cdot ^\circ C}$$

$$c = \underline{2500 \frac{J}{kg \cdot ^\circ C}}$$

$$= \frac{30000 J}{5 kg \cdot 24^\circ C}$$

Thermodynamics Study Guide Answers

Arshad Iqbal



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Study Guide/Solutions Manual for Organic Chemistry K. Peter C. Vollhardt, Neil E. Schore, 2018-01-12 Updated for the Eighth Edition of Vollhardt Schore Organic Chemistry and written by the book's coauthor Neil Schore this invaluable manual includes chapter introductions that highlight new material chapter outlines detailed comments for each chapter section a glossary and solutions to the end of chapter problems presented in a way that shows students how to reason their way to the answer

Thermodynamics Ibrahim Dincer, 2020-08-18 Presents a unique stepwise exergy based approach to thermodynamic concepts systems and applications Thermodynamics A Smart Approach redefines this crucial branch of engineering as the science of energy and exergy rather than the science of energy and entropy to provide an innovative step by step approach for teaching understanding and practicing thermodynamics in a clearer and easier way Focusing primarily on the concepts and balance equations this innovative textbook covers exergy under the second law of thermodynamics discusses exergy matters and relates thermodynamics to environmental impact and sustainable development in a clear simple and understandable manner It aims to change the way thermodynamics is taught and practiced and help overcome the fear of thermodynamics Author Ibrahim Dincer a pioneer in the areas of thermodynamics and sustainable energy technologies draws upon his multiple decades of experience teaching and researching thermodynamics to offer a unique exergy based approach to the subject Enabling readers to easily comprehend and apply thermodynamic principles the text organizes thermodynamics into seven critical steps property state process cycle first law of thermodynamics second law of thermodynamics and performance assessment and provides extended teaching tools for systems and applications Precise student friendly chapters cover fundamental concepts thermodynamic laws conventional and innovative power and refrigeration cycles and more This textbook Covers a unique approach in teaching design analysis and assessment of thermodynamic systems Provides lots of examples for every subject for students and instructors Contains hundreds of illustrations figures and tables to better illustrate contents Includes many conceptual questions and study problems Features numerous systems related examples and practical applications Thermodynamics A Smart Approach is an ideal textbook for undergraduate students and graduate students of engineering and applied science as well researchers scientists and practicing engineers seeking a precise and concise textbook and or reference work

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Katherine J. Denniston, 1995 This student study guide has been written to accompany Foundations of Inorganic Organic and Biological Chemistry by Carey Denniston and Topping It was designed to complement the text not to be used in place of the text Each chapter of the study guide contains the following sections reorganized and expanded set of learning goals concise chapter summary in chapter solved problems list of key terms review problems Preface

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properties while maintaining a solid treatment of chemical principles Common chemicals are used whenever possible as examples in both theoretical discussions and in problems and exercises Incorporates many pedagogical aids each chapter begins with a brief table of contents and each section begins with a preview of topics covered Chapters include frequent margin comments figures and photographs

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