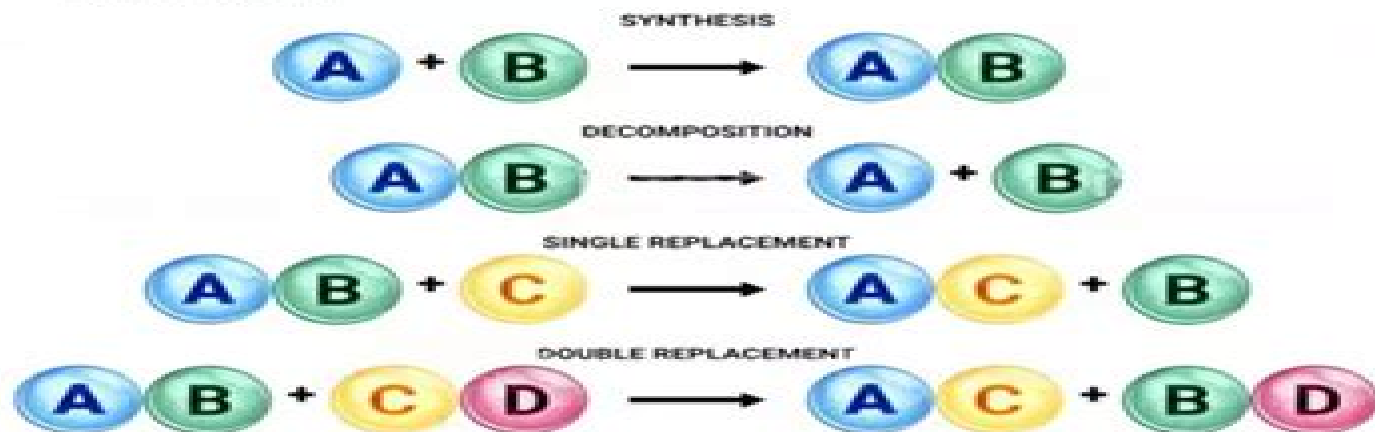


III. Types of Chemical Reactions

The 4 primary types of chemical reactions are:

1. **Combination (synthesis) reaction:** Two or more reactants combine to make a new product. Going from simple to complex molecules.
2. **Decomposition reaction:** A single reactant breaks down to form two or more products. Going from complex to simple molecules.
3. **Single Displacement reaction:** A single element replaces a similar element of an adjacent reactant compound.
4. **Double Displacement reaction:** Two compounds exchange elements, producing two new compounds.



ACTIVITY

Identify the type of reaction for the following equations.

Chemical Reaction	Type of reaction
$\text{Cu}_{(s)} + 2\text{AgNO}_{3(aq)} \rightarrow \text{Cu}(\text{NO}_3)_{2(aq)} + 2\text{Ag}_{(s)}$	
$2\text{KI}_{(aq)} + \text{Pb}(\text{NO}_3)_{2(aq)} \rightarrow 2\text{KNO}_{3(aq)} + \text{PbI}_{2(s)}$	
$2\text{H}_2\text{O}_{2(l)} \rightarrow 2\text{H}_2\text{O}_{(l)} + \text{O}_{2(g)}$	
$\text{C}_6\text{H}_{12}\text{O}_6 \rightarrow 2\text{C}_2\text{H}_5\text{OH} + 2\text{CO}_2$	
$\text{Mg}_{(s)} + \text{CuSO}_{4(aq)} \rightarrow \text{MgSO}_{4(aq)} + \text{Cu}_{(s)}$	
$2\text{C}_8\text{H}_{18} + 25\text{O}_2 \rightarrow 16\text{CO}_2 + 18\text{H}_2\text{O}$	

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