

Marking Scheme

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Page 1
Section A

1. (a) (i)



(1)

(ii) Same test

(1)

Carbon compounds give a brick-red flame.

(1)

(iii) (i) Isotopes are atoms of the same element with same proton number but different neutron numbers.

(1)

	Number of protons	Number of neutrons
$^{65}_{28}\text{Ni}$	28	37

(1)

(ii) (i) They have the same number of electrons in their outermost shells.

(1)

(ii) It has similar chemical properties to Cu since, they can replace copper of the Cu required.

(1)

2. (a) (i) $2\text{AgCl} \rightarrow 2\text{Ag} + \text{Cl}_2$

(1)

(ii) The oxidation no. of Ag decreases and the oxidation no. of Cl increases.

(1)

(iii) No. of moles of $\text{AgCl} = 0.10 \text{ mol}$ ($0.037 \text{ g} \div 37$)

No. of moles of $\text{Ag} = 2 \times$ No. of moles of AgCl

Mass of Ag that can be obtained = $107.9 \times$ No. of moles of Ag = $\frac{0.037 \times 216}{107.9} = 0.74$

(1)

(iv) The black oxide changes to metallic brown metal

(1)

(v) The metal obtained can conduct electricity.

(1)

(vi) $\text{CuO} + \text{H}_2 \rightarrow \text{Cu} + \text{H}_2\text{O}$

(1)

(vii) Hydrogen is explosive / flammable.

(1)

(b) No. The reactivity of Cu and Ag can only be compared using the same reaction.

(1)

22

3. (a) (i) Sodium hydroxide is very caustic.

(1)

(ii) The sodium hydroxide is used.

(1)

(b) (i) Copper cannot displace H^+ (aq) from HCl (aq).

(1)

(ii) Add Zn/Mg from HCl (aq).

(1)

(c) (i) When water is added to concentrated H_2SO_4 , a lot of heat is produced. The heat can cause spilling out of the concentrated acid solution.

(1)

(ii) Add concentrated H_2SO_4 to water slowly and stir the mixture.

(1)

4. (a) (i) Chlorinate containing substances are harmful to marine life / water pollution.

(1)

(ii) (i) $\text{SO}_4^{2-} + \text{H}_2\text{O} \rightarrow \text{HSO}_4^- + \text{OH}^- + \text{H}^+$

(1)

(ii) $\text{CO}_3^{2-} + 2\text{H}^+ + \text{H}_2\text{O} \rightarrow 2\text{CO}_2 + \text{H}_2\text{O}$

(1)

(iii) $\text{CO}_3^{2-} + \text{H}^+ + \text{H}_2\text{O} \rightarrow \text{CO}_2 + \text{H}_2\text{O} + \text{HCO}_3^-$

(1)

(iii) $\text{HCO}_3^- + \text{H}_2\text{O} + \text{H}^+ \rightarrow \text{H}_2\text{CO}_3 + \text{H}_2\text{O}$

(1)

5. (a) Any TWO of the following points:

(1)

- Substances and series can be represented by a semi-empirical formula.

The general formula for pentane and octane is $\text{C}_n\text{H}_{2n+2}$. Adjacent members differ by one $-\text{CH}_2-$.

- There are gradual changes in physical properties among the members of a homologous series. The boiling point / melting point / density / refractive index is higher than that of pentane.

- Members of the same homologous series have similar chemical properties.

Both pentane and octane can undergo substitution reactions with Br_2/CCl_4 .

(b) Octane, C_8H_{18} has a higher percentage of carbon by mass. Its chance to undergo incomplete combustion is given carbon is higher.

(1)

(c) Any TWO of the following:

(1)

$\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$

$\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$

$\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$

23

6. (a) (i)



(1)

(ii) It is a mixture of polymeric molecules of different chain lengths.

(1)

(iii) (i) condenser

(1)

(ii) Electric heating mantle / oil bath / sand bath should be used because benzene is flammable.

(1)

iii. Draw the solution in a fume cupboard because styrene vapour is irritant.

(1)

(iv) addition polymerisation

(1)

(v) (i) Electrically heating can be prevented.

(1)

(ii) Plastic does not conduct readily.

(1)

(iii) Low density

(1)

7. Chemical Knowledge

(1)

Insert two of the metal strips into the beaker to form a chemical cell. Measure the voltage of the cell using the multimeter. Complete the electric circuit.

Control variables in the experiment such as:

- the size of the strips
- the separation / distance between strips should be the same in each trial
- the temperature

The highest voltage can be obtained using a magnesium strip and a copper strip. (It is because among the three metals, Mg occupies the highest position in the e.m.f. and Cu the lowest position).

Adjust the distance / position between the two metal strips until the maximum voltage / current is obtained.

Apparatus / Concentration

(1)

Section B

8. (a)

$$\begin{aligned} \text{mole ratio} &= \frac{\text{Fe}}{\text{Fe}_2\text{O}_3} = \frac{0.8}{1.2} \\ &= 0.667 = 0.667 \times \frac{2}{3} \\ &= \frac{2}{3} \end{aligned}$$

Empirical formula of X is Fe_2O_3

(b) Let mole ratio of FeO to Fe_2O_3 be $x : y$

$$\frac{\text{mole of Fe}}{\text{mole of O}} = \frac{x + 2y}{x + y} = \frac{1}{2}$$

(1)

X is a mixture of FeO and Fe_2O_3 in a mole ratio of 2 : 1.
or, X is not a mixture. In X, two-third of the total oxide is an oxidation number +2, while one-third is an oxidation number +3.

(1)

9. (a) (i) Hydrogen

(1)

(ii) manufacture of H_2SO_4 / manufacture of CH_3COH / manufacture of hydrochloric acid / fertilising of vegetable soils

(1)

(b) (i) Chlorine

(1)

(ii) At very low concentrations, Cl_2 will be formed + HCl (aq) discharged.

(1)

Reaction of OH^- in HCl is higher than that of Cl^- / OH^- is a stronger reducing agent than Cl^- / OH^- loses electron more readily than Cl^- .

(1)

(c) It is a reducing agent / It loses electrons

(1)

MnO_2 is an oxidising agent / MnO_2 gains electrons

(1)

When the cell is connected in an external circuit, electrons will flow through the external circuit.

(1)

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**Peter Blake, Elfed Charles, Kathryn
Foster**



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 - The Rise of Digital Reading Wjec Chemistry Marking Scheme 2005
 - Advantages of eBooks Over Traditional Books
2. Identifying Wjec Chemistry Marking Scheme 2005
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Wjec Chemistry Marking Scheme 2005
 - User-Friendly Interface
4. Exploring eBook Recommendations from Wjec Chemistry Marking Scheme 2005

- Personalized Recommendations
- Wjec Chemistry Marking Scheme 2005 User Reviews and Ratings
- Wjec Chemistry Marking Scheme 2005 and Bestseller Lists
- 5. Accessing Wjec Chemistry Marking Scheme 2005 Free and Paid eBooks
 - Wjec Chemistry Marking Scheme 2005 Public Domain eBooks
 - Wjec Chemistry Marking Scheme 2005 eBook Subscription Services
 - Wjec Chemistry Marking Scheme 2005 Budget-Friendly Options
- 6. Navigating Wjec Chemistry Marking Scheme 2005 eBook Formats
 - ePub, PDF, MOBI, and More
 - Wjec Chemistry Marking Scheme 2005 Compatibility with Devices
 - Wjec Chemistry Marking Scheme 2005 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Wjec Chemistry Marking Scheme 2005
 - Highlighting and Note-Taking Wjec Chemistry Marking Scheme 2005
 - Interactive Elements Wjec Chemistry Marking Scheme 2005
- 8. Staying Engaged with Wjec Chemistry Marking Scheme 2005
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Wjec Chemistry Marking Scheme 2005
- 9. Balancing eBooks and Physical Books Wjec Chemistry Marking Scheme 2005
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Wjec Chemistry Marking Scheme 2005
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Wjec Chemistry Marking Scheme 2005
 - Setting Reading Goals Wjec Chemistry Marking Scheme 2005
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Wjec Chemistry Marking Scheme 2005

- Fact-Checking eBook Content of Wjec Chemistry Marking Scheme 2005
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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