

3.2.10 aromatic chemistry

structure of benzene

- benzene formula: C_6H_6
- planar cyclic molecule and non-polar
- each carbon is bonded to two other carbons in a ring and one hydrogen
- the bond lengths between the carbon atoms in benzene are equal, and hence the length of the ring
- π -electrons are made up of the carbon atom orbitals to form a ring of delocalised electron (delocalised system) or resonance
- all the bonds in the ring are the same length. $\rightarrow$ trigonal planar which is an indication of the $C-C$ (134pm) and $C-H$ (109pm)



ring of delocalised electron can be represented by a circle in the ring of carbon-carbon bonds (circle model)

delocalised π system means benzene more stable



stability of benzene

benzene's unusual structure: Chemical systems $\rightarrow$ 1,3,5-triene



benzene proposed that benzene occurred in a ring and single bonds

predicted with the structure:

- benzene should undergo addition reactions
- benzene is "aromatic" more stable than benzene's structural isomers $\rightarrow$ 1,3,5-triene
- bond lengths: $C-C$ (134pm) $C-H$ (109pm)
- in benzene all bonds are 139pm

when comparing enthalpy change of hydrogenation for benzene with the enthalpy change of hydrogenation for 1,3,5-triene we see that benzene is not more stable



hydrogenation for 1,3,5-triene had
this is $\rightarrow$ hydrogenated the structure change is $\rightarrow$ "aromatic"



if benzene had three double bonds
expected enthalpy of hydrogenation would be $\rightarrow$ "aromatic"



benzene has experimental enthalpy of hydrogenation of benzene is $\rightarrow$ "aromatic"

benzene has more stable than 1,3,5-triene
this is because of the delocalized pi system

aromaticity criteria

benzene: aromatic molecule in benzene ring

when benzene is the only functional group
prefix is functional group: methyl $\rightarrow$ benzene



when benzene is not the only functional group
prefix is functional group: methyl $\rightarrow$ benzene



benzene has aromaticity

benzene has aromaticity because it has a delocalized pi system
it has a ring of delocalized pi electrons

aromaticity

benzene has 6 π electrons

benzene has 6 π electrons: delocalized system $\rightarrow$ aromaticity
 $\rightarrow$ 6 π electrons

benzene has 6 π electrons: delocalized system $\rightarrow$ 6 π electrons

benzene has 6 π electrons: delocalized system

benzene has 6 π electrons: delocalized system $\rightarrow$ aromaticity

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