

Day 1

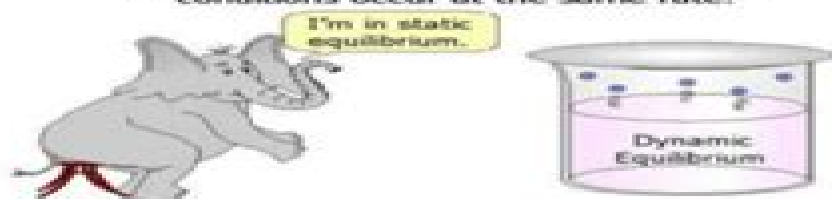
A. Recognizing Equilibrium (pg 424)

I. Gaseous Equilibrium : Consider the reaction : $\text{H}_2 (\text{g}) + \text{I}_2 (\text{g}) \rightleftharpoons 2 \text{HI} (\text{g})$

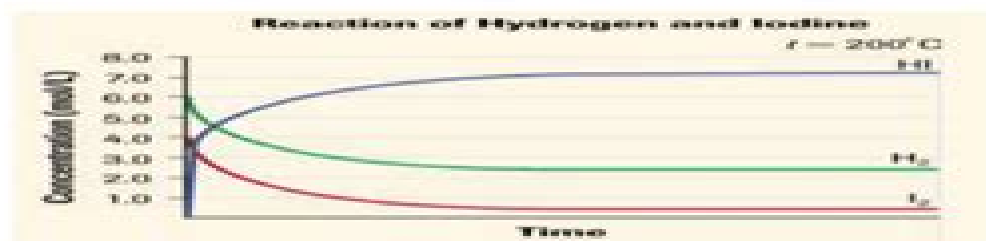
6.00 mole of H_2 and 4.00 mole of I_2 are placed in a 1.00 L vessel ; some of the H_2 and I_2 molecules collide and react forming HI molecules ; as the HI molecules accumulate they begin reacting with each other forming H_2 and I_2 . At first the rate of the forward reaction is _____ than the rate of the reverse reaction but as time proceeds the forward reaction _____ in rate and the reverse _____ in rate. At some point in time the rate of the forward and reverse reaction are equal and there appears to be no reaction going on. The above reaction is easy to follow since I_2 is violet in colour and HI is colourless. At first the colour of the system is _____ and the colour will _____ as the reaction proceeds. When there is no further change in colour we know the system is in _____. Referred to as _____ equilibrium since reactions are still going on but in a balanced state.

Equilibrium

A state where the forward and reverse conditions occur at the same rate.



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[initial][i]	_____	_____	_____
[change][c]	_____	_____	_____
[equilibrium][e]	_____	_____	_____

Percent reaction

- Similar to percent yield
- The yield of product measured **at equilibrium** compared with the maximum possible yield of product

Unit 4 Chemical Equilibrium

Educart

A decorative red circular graphic with a gradient, appearing as a semi-circle or a stylized arrow pointing to the right, located on the right side of the light blue horizontal bar.

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